

Question 1

1(a)iii) 247 g per 100 cm³.

1(a)(iv) Mass of KI dissolved in 100 cm³ at 70 °C = 247 g.
Mass of KI dissolved in 100 cm³ at 30 °C = 180 g.

Mass of KI that would be precipitated out from solution on cooling from 70 °C to 30 °C
= (247 – 180) g = 67 g.

1(a)(v) Molar mass of KI = $(127 + 39) \text{ g mol}^{-1} = 166 \text{ g mol}^{-1}$.
 Moles of KI in 180 g KI = $(180/166) \text{ mol} = 1.08 \text{ mol}$.
 Hence moles of KI in 100 cm^3 of saturated solution at $30^\circ\text{C} = 1.08 \text{ mol}$.
 Moles of KI in $1000 \text{ cm}^3 = (1.08 \times 10) = 10.8 \text{ mol}$.
 Therefore concentration of saturated solution of KI at $30^\circ\text{C} = 10.8 \text{ mol dm}^{-3}$.

1(b) KI is an ionic solid in which cations and anions are held together by ionic bonds. Water is a polar covalent solvent so it can cause the ions in KI to dissociate. Ethanol is also a polar covalent solvent but it is not as polar as water.

1(c)(i) Distilled water, stirring rod, beakers, aqueous barium chloride solution, filter funnel, filter paper, Bunsen burner, tripod stand, wire gauze and evaporating dish.

1. Dissolve the solid mixture in distilled water.
2. Add aqueous barium chloride solution slowly with stirring until no more precipitate forms.
3. Filter of the precipitate of barium sulfate.
4. Evaporate the filtrate to obtain solid NaCl.

1(c)(iii)

- Colourless solution forms when salt mixture dissolves.
- White precipitate forms when aqueous barium chloride is added.
- Colourless filtrate is collected upon filtration.
- White solid forms upon evaporation of filtrate.

1(d)(i) Inference

- NO_2 gas formed.
- Acidic gas present: NO_3^- ions present.

1(d)(ii) Inference

- Pb^{2+} ions present.
- $\text{Pb}^{2+}_{(aq)} + \text{I}^{-}_{(aq)} \rightarrow \text{PbI}_{2(s)}$

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