

Name: \_\_\_\_\_

## Lab: Electron Arrangements

**Objectives:** You will write short-handled electron configurations, draw orbital filling diagrams, and write electron-dot diagrams.

**Procedure:** Complete the table below.

Element	Atomic Number	Short-handled Electron Configuration	Orbital Filling Diagrams (Use Box Symbols)	Electron-dot Diagram
Lithium	3	$1s^2 2s^1$	$1s^2 \boxed{\uparrow\downarrow}$	$\cdot\cdot$
Boron	5	$1s^2 2s^2 2p^1$		
Carbon	6	$1s^2 2s^2 2p^2$		
Nitrogen	7	$1s^2 2s^2 2p^3$		
Oxygen	8	$1s^2 2s^2 2p^4$		
Fluorine	9	$1s^2 2s^2 2p^5$		
Neon	10	$1s^2 2s^2 2p^6$		
Sodium	11	$1s^2 2s^2 2p^6 3s^1$		
Magnesium	12	$1s^2 2s^2 2p^6 3s^2$		
Aluminum	13	$1s^2 2s^2 2p^6 3s^2 3p^1$		
Silicon	14	$1s^2 2s^2 2p^6 3s^2 3p^2$		

# Name Electron Arrangements Ws Answers

**Anthony S. Fauci**



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