

# Names & Formulas of Compounds

| Positive Ions (cations) |                  | Negative Ions (anions) |                    |
|-------------------------|------------------|------------------------|--------------------|
| Name                    | Formula          | Name                   | Formula            |
| Hydrogen                | $\text{H}^+$     | Chloride               | $\text{Cl}^-$      |
| Sodium                  | $\text{Na}^+$    | Bromide                | $\text{Br}^-$      |
| Silver                  | $\text{Ag}^+$    | Fluoride               | $\text{F}^-$       |
| Potassium               | $\text{K}^+$     | Iodide                 | $\text{I}^-$       |
| Lithium                 | $\text{Li}^+$    | Hydroxide              | $\text{OH}^-$      |
| Ammonium                | $\text{NH}_4^+$  | Nitrate                | $\text{NO}_3^-$    |
| Barium                  | $\text{Ba}^{2+}$ | Oxide                  | $\text{O}^{2-}$    |
| Calcium                 | $\text{Ca}^{2+}$ | Sulphide               | $\text{S}^{2-}$    |
| Copper(II)              | $\text{Cu}^{2+}$ | Sulphate               | $\text{SO}_4^{2-}$ |
| Magnesium               | $\text{Mg}^{2+}$ | Carbonate              | $\text{CO}_3^{2-}$ |
| Zinc                    | $\text{Zn}^{2+}$ | Hydrogencarbonate      | $\text{HCO}_3^-$   |
| Lead                    | $\text{Pb}^{2+}$ |                        |                    |
| Iron(II)                | $\text{Fe}^{2+}$ |                        |                    |
| Iron(III)               | $\text{Fe}^{3+}$ |                        |                    |
| Aluminium               | $\text{Al}^{3+}$ |                        |                    |

# Names And Formulas For Compounds

**Ralph A. Burns, John William Hill**



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