

## WORKSHEET – Naming Compounds KEY

### Part I – Naming Compounds and Numbering Atoms

FORMULA	NAME	# of DIFFERENT ATOMS	TOTAL # OF ATOMS
1. $\text{CaCl}_2$	Calcium chloride	1 Ca; 2 Cl	3
2. $\text{AgCl}$	<i>Silver chloride</i>	1 Ag; 1 Cl	2
3. $\text{MgO}$	<i>Magnesium oxide</i>	1 Mg; 1 O	2
4. $\text{NaBr}$	<i>Sodium bromide</i>	1 Na; 1 Br	2
5. $\text{Al}_2\text{O}_3$	<i>Aluminum oxide</i>	2 Al; 3 O	5
6. $\text{KI}$	<i>Potassium iodide</i>	1 K; 1 I	2
7. $\text{ZnCl}_2$	<i>Zinc chloride</i>	1 Zn; 2 Cl	3
8. $\text{Li}_2\text{O}$	<i>Lithium oxide</i>	2 Li; 1 O	3
9. $\text{BaO}$	<i>Barium oxide</i>	1 Ba; 1 O	2
10. $\text{CaBr}_2$	<i>Calcium bromide</i>	1 Ca; 2 Br	3
11. $\text{MgCl}_2$	<i>Magnesium chloride</i>	1 Mg; 2 Cl	3
12. $\text{AgI}$	<i>Silver iodide</i>	1 Ag; 1 I	2
13. $\text{ZnS}$	<i>Zinc sulphide</i>	1 Zn; 1 S	2
14. $\text{BaF}_2$	<i>Barium fluoride</i>	1 Ba; 2 F	3
15. $\text{Ca}_3\text{P}_2$	<i>Calcium phosphide</i>	3 Ca; 2 P	5
16. $\text{Na}_2\text{O}$	<i>Sodium oxide</i>	2 Na; 1 O	3
17. $\text{AlN}$	<i>Aluminum nitride</i>	1 Al; 1 N	2
18. $\text{K}_3\text{N}$	<i>Potassium nitride</i>	3 K; 1 N	4
19. $\text{Ca}_2\text{C}$	<i>Calcium carbide</i>	2 Ca; 1 C	3
20. $\text{K}_2\text{O}$	<i>Potassium oxide</i>	2 K; 1 O	3

# Naming Compounds Answer Key

**James E. Banks**

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