

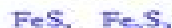
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Why?

When working in chemistry, it is often convenient to write a chemical in symbols. For example we might write down the substance table salt as NaCl . In talking about chemistry however, it is a bit tacky to say "en-ay see-ell" when we want to refer to a substance. Also, in formal writing we should use the name of the compound rather than its symbols. Therefore we need to learn how to say the proper names of ionic substances.

1	H ⁺									
2	Li ⁺	He ²⁺					N ³⁻	O ²⁻	F ⁻	
3	Na ⁺	Mg ²⁺	Transition elements				Al ³⁻	P ³⁻	S ²⁻	Cl ⁻
4	K ⁺	Ca ²⁺	Fe ²⁺ Fe ³⁺	Ni ²⁺ Ni ³⁺	Cu ⁺ Cu ²⁺	Zn ²⁺				Br ⁻
5	Rb ⁺	Sr ²⁺			Ag ⁺			Sn ²⁺ Sn ⁴⁺		I ⁻
6		Ba ²⁺				Hg ₂ ²⁺ Hg ²⁺		Pb ²⁺ Pb ⁴⁺		
	Cations						Anions			

1. Consider the ions of potassium (K) and sulfur (S). Write the chemical formulas for the ionic compound involving these ions, using the simplest ratio(s) of potassium (K) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.
 K_2S
2. Consider the ions of iron (Fe) and sulfur (S). Write chemical formulas for all possible ionic compounds involving these ions, using the simplest ratio(s) of iron (Fe) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.



NaCl	Sodium chloride	Zn_3P_2	Zinc phosphide
CaS	Calcium sulfide	Al_2O_3	Aluminium oxide
Ag_2S	Silver sulfide	SnCl_4	Stannic chloride

3. Circle the symbol for the metal in each of the compounds in Model 2. If doing electronically, you can just **BOLD** the symbol. Remember your PT!

Naming Ionic Compounds Key

RS Peters

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'RS Peters' text.

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