

Naming and Equations Review

Name the following compounds. Remember, they may be either ionic or covalent compounds, so make sure you use the right naming method!

- 1) LiBr _____
- 2) Na_2O _____
- 3) CaSO_4 _____
- 4) $\text{Be}(\text{NO}_3)_2$ _____
- 5) NO _____

Write the formulas for the following compounds. Remember, they may be either ionic or covalent compounds, so make sure you use the right method!

- 6) sodium nitride _____
- 7) sulfur dioxide _____
- 8) ammonia _____
- 9) calcium phosphate _____
- 10) aluminum nitrate _____

Balance the following equations:

- 11) _____ $\text{NaCN} +$ _____ $\text{CuCO}_3 \rightarrow$ _____ $\text{Na}_2\text{CO}_3 +$ _____ $\text{Cu}(\text{CN})_2$
- 12) _____ $\text{Al}(\text{OH})_3 +$ _____ $\text{Cs} \rightarrow$ _____ $\text{CsOH} +$ _____ Al
- 13) _____ $\text{MgS} +$ _____ $\text{OF}_2 \rightarrow$ _____ $\text{MgF}_2 +$ _____ SO
- 14) _____ $\text{Se} +$ _____ $\text{H}_2 \rightarrow$ _____ H_2S
- 15) _____ $\text{LiCl} +$ _____ $\text{Br}_2 \rightarrow$ _____ $\text{LiBr} +$ _____ Cl_2
- 16) _____ $\text{Pb}(\text{NO}_3)_2 +$ _____ $\text{HF} \rightarrow$ _____ $\text{HNO}_3 +$ _____ PbF_2
- 17) _____ $\text{GaBr}_3 +$ _____ $\text{Na}_2\text{SO}_4 \rightarrow$ _____ $\text{NaBr} +$ _____ $\text{Ga}_2(\text{SO}_4)_3$
- 18) _____ $\text{P}_2\text{S}_2 \rightarrow$ _____ $\text{P}_4 +$ _____ S_8

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Daniel F McAuley



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