

Limiting Reactants Practice



a) What is the limiting reactant if 0.866 mol of NO is mixed with 0.503 mol of O₂ to create NO₂?

$$0.866 \text{ mol NO} \times \frac{2 \text{ mol NO}_2}{2 \text{ mol NO}} = 0.866 \text{ mol NO}_2$$

limiting reactant product made

$$0.503 \text{ mol O}_2 \times \frac{2 \text{ mol NO}_2}{1 \text{ mol O}_2} = 1.006 \text{ mol NO}_2$$

b) Calculate the number of moles of NO₂ produced. (Answer should already be solved from the previous questions)

0.866 mol NO₂



a) If 7.40 g of O₃ reacts with 0.670 g of NO, which compound will be the limiting reactant?

O₃
O: 3 × 15.999 = 47.997 g/mol
⇒ 1 mol O₃ = 47.997 g O₃

$$7.40 \text{ g O}_3 \times \frac{1 \text{ mol O}_3}{47.997 \text{ g O}_3} \times \frac{1 \text{ mol NO}_2}{1 \text{ mol O}_3} \times \frac{46.005 \text{ g NO}_2}{1 \text{ mol NO}_2} = 7.093 \text{ g NO}_2$$

NO
N: 1 × 14.007 = 14.007
O: 1 × 15.999 = 15.999
⇒ 1 mol NO = 30.006 g/mol

$$0.670 \text{ g NO} \times \frac{1 \text{ mol NO}}{30.006 \text{ g NO}} \times \frac{1 \text{ mol NO}_2}{1 \text{ mol NO}} \times \frac{46.005 \text{ g NO}_2}{1 \text{ mol NO}_2} = 1.027 \text{ g NO}_2$$

limiting reactant product made

b) How many grams of NO₂ will be produced? (Answer should already be solved from the previous questions)

1.027 g NO₂

NO₂
N: 1 × 14.007 = 14.007
O: 2 × 15.999 = 31.998
⇒ 1 mol NO₂ = 46.005 g/mol

c) Calculate mass, in grams, of the excess reactant remaining at the end of the reaction. How many moles of excess reactant would this be?

$$0.670 \text{ g NO} \times \frac{1 \text{ mol NO}}{30.006 \text{ g NO}} \times \frac{1 \text{ mol O}_3}{1 \text{ mol NO}} \times \frac{47.997 \text{ g O}_3}{1 \text{ mol O}_3} = 1.072 \text{ g O}_3 \text{ used!}$$

$$\underbrace{7.40 \text{ g O}_3}_{\text{original}} - \underbrace{1.072 \text{ g O}_3}_{\text{used}} = \underline{6.328 \text{ g O}_3 \text{ leftover}}$$

Limiting Reactants Study Guide Answer Key

Paul B. Kelter, Steven S. Zumdahl



Limiting Reactants Study Guide Answer Key:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03

The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

Chemistry, Student Study Guide James E. Brady, Fred

Senese, 2008-01-28. The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst. Carbon nanotubes are a recently discovered allotrope of carbon. Three other allotropes of carbon, buckyballs, graphite, and diamond, are illustrated at the left, as is the molecule methane, CH_4 , from which nanotubes and buckyballs can be made. The element carbon forms an amazing number of compounds with structures that follow from simple methane, found in natural gas, to the complex macromolecules that serve as the basis of life on our planet. The study of chemistry also follows from the simple to the more complex, and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement.

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Kelter, Steven S. Zumdahl, 2004-04 **Study Guide General Chemistry, Second Edition, Becker/Wentworth**
 Becker, Benjamin L. Carroll, 1980 Prentice Hall Chemistry, 2000 Applied Chemistry Siddharth Venkatesh, 2025-02-20

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Table of Contents Limiting Reactants Study Guide Answer Key

1. Understanding the eBook Limiting Reactants Study Guide Answer Key
 - The Rise of Digital Reading Limiting Reactants Study Guide Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reactants Study Guide Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting Reactants Study Guide Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reactants Study Guide Answer Key
 - Personalized Recommendations
 - Limiting Reactants Study Guide Answer Key User Reviews and Ratings
 - Limiting Reactants Study Guide Answer Key and Bestseller Lists
5. Accessing Limiting Reactants Study Guide Answer Key Free and Paid eBooks
 - Limiting Reactants Study Guide Answer Key Public Domain eBooks
 - Limiting Reactants Study Guide Answer Key eBook Subscription Services
 - Limiting Reactants Study Guide Answer Key Budget-Friendly Options
6. Navigating Limiting Reactants Study Guide Answer Key eBook Formats

- ePub, PDF, MOBI, and More
 - Limiting Reactants Study Guide Answer Key Compatibility with Devices
 - Limiting Reactants Study Guide Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Limiting Reactants Study Guide Answer Key
 - Highlighting and Note-Taking Limiting Reactants Study Guide Answer Key
 - Interactive Elements Limiting Reactants Study Guide Answer Key
8. Staying Engaged with Limiting Reactants Study Guide Answer Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting Reactants Study Guide Answer Key
9. Balancing eBooks and Physical Books Limiting Reactants Study Guide Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reactants Study Guide Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Limiting Reactants Study Guide Answer Key
- Setting Reading Goals Limiting Reactants Study Guide Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Limiting Reactants Study Guide Answer Key
- Fact-Checking eBook Content of Limiting Reactants Study Guide Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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