

**12 • The Gas Laws****BOYLE'S LAW**

Boyle's Law states that the volume of a gas varies inversely with its pressure if temperature is held constant. (If one goes up, the other goes down.) We use the formula:

$$P_1 \times V_1 = P_2 \times V_2$$

Solve the following problems (assuming constant temperature). Assume all number are 3 significant figures.

1. A sample of oxygen gas occupies a volume of 250 mL at 740 torr pressure. What volume will it occupy at 800 torr pressure?

$$P_1 V_1 = P_2 V_2$$

$$V_2 = \frac{P_1 V_1}{P_2} = \frac{(740 \text{ torr})(250 \text{ mL})}{(800 \text{ torr})} = \boxed{231 \text{ mL}}$$

2. A sample of carbon dioxide occupies a volume of 3.50 Liters at 125 kPa pressure. What pressure would the gas exert if the volume was decreased to 2.00 liters?

$$P_2 = \frac{P_1 V_1}{V_2} = \frac{(125 \text{ kPa})(3.50 \text{ L})}{(2.00 \text{ L})} = \boxed{219 \text{ kPa}}$$

3. A 2.00-Liter container of nitrogen had a pressure of 3.20 atm. What volume would be necessary to decrease the pressure to 1.00 atm?

$$P_1 = 3.20 \text{ atm} \quad V_1 = 2.00 \text{ L}$$

$$P_2 = 1.00 \text{ atm}$$

$$V_2 = \frac{P_1 V_1}{P_2} = \frac{(3.20 \text{ atm})(2.00 \text{ L})}{(1.00 \text{ atm})} = \boxed{6.40 \text{ L}}$$

4. Ammonia gas occupies a volume of 450 mL as a pressure of 720 mmHg. What volume will it occupy at standard pressure (760 mmHg)?

$$P_1 = 720 \text{ mmHg} \quad V_1 = 450 \text{ mL}$$

$$P_2 = 760 \text{ mmHg}$$

$$V_2 = \frac{P_1 V_1}{P_2} = \frac{(720 \text{ mmHg})(450 \text{ mL})}{(760 \text{ mmHg})} = \boxed{426 \text{ mL}}$$

5. A 175 mL sample of neon had its pressure changed from 75.0 kPa to 150 kPa. What is its new volume?

$$V_1$$

$$V_2 = \frac{P_1 V_1}{P_2} = \frac{(75.0 \text{ kPa})(175 \text{ mL})}{150 \text{ kPa}} = \boxed{87.5 \text{ mL}}$$

6. A sample of hydrogen at 1.50 atm had its pressure decreased to 0.50 atm producing a new volume of 750 mL. What was the sample's original volume?

$$V_1 = \frac{P_2 V_2}{P_1} = \frac{(0.50 \text{ atm})(750 \text{ mL})}{(1.50 \text{ atm})} = \boxed{250. \text{ mL}}$$

7. Chlorine gas occupies a volume of 1.20 liters at 720 torr pressure. What volume will it occupy at 1 atm pressure?

$$V_2 = \frac{P_1 V_1}{P_2} = \frac{(720 \text{ torr})(1.20 \text{ L})}{(1 \text{ atm} \times \frac{760 \text{ torr}}{1 \text{ atm}})} = \boxed{1.14 \text{ L}}$$

8. Fluorine gas exerts a pressure of 900 torr. When the pressure is changed to 1.50 atm, its volume is 250 mL. What was the original volume?

$$V_1 = \frac{P_2 V_2}{P_1} = \frac{(1.50 \text{ atm})(250 \text{ mL})}{(900 \text{ torr} \times \frac{1 \text{ atm}}{760 \text{ torr}})} = \boxed{317 \text{ mL}}$$

# J Weston Walch Gas Law Problems Answers

**M Tight**



## **J Weston Walch Gas Law Problems Answers:**

## Unveiling the Energy of Verbal Beauty: An Mental Sojourn through **J Weston Walch Gas Law Problems Answers**

In a world inundated with displays and the cacophony of instant communication, the profound energy and mental resonance of verbal beauty frequently diminish in to obscurity, eclipsed by the regular onslaught of noise and distractions. Yet, set within the lyrical pages of **J Weston Walch Gas Law Problems Answers**, a fascinating work of literary brilliance that impulses with raw feelings, lies an remarkable trip waiting to be embarked upon. Penned by a virtuoso wordsmith, this exciting opus instructions viewers on an emotional odyssey, delicately revealing the latent possible and profound affect embedded within the delicate internet of language. Within the heart-wrenching expanse with this evocative analysis, we will embark upon an introspective exploration of the book is central subjects, dissect their captivating publishing design, and immerse ourselves in the indelible impact it leaves upon the depths of readers souls.

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