

## Limiting Reactants Worksheet KEY

Answer the following questions, using stoichiometry to convert between moles of one substance and moles of another substance within a reaction and using molar masses to convert between grams of a substance and moles of that substance.

Given the following reaction:  $\text{C}_5\text{H}_{12} + 8 \text{O}_2 \rightarrow 5 \text{CO}_2 + 6 \text{H}_2\text{O}$

1. If 5.25 grams of pentane ( $\text{C}_5\text{H}_{12}$ ) are completely reacted with oxygen ( $\text{O}_2$ ), how many grams of carbon dioxide ( $\text{CO}_2$ ) will be produced?

$$5.25 \text{ g C}_5\text{H}_{12} \times \frac{1 \text{ mol C}_5\text{H}_{12}}{72.151 \text{ g C}_5\text{H}_{12}} = 0.07276 \text{ mol C}_5\text{H}_{12}$$

$$0.07276 \text{ mol C}_5\text{H}_{12} \times \frac{5 \text{ mol CO}_2}{1 \text{ mol C}_5\text{H}_{12}} = 0.3638 \text{ mol CO}_2$$

$$0.3638 \text{ mol CO}_2 \times \frac{44.009 \text{ g CO}_2}{1 \text{ mol CO}_2} = 16.0 \text{ g CO}_2$$

2. If 9.75 grams of oxygen ( $\text{O}_2$ ) are completely reacted with pentane ( $\text{C}_5\text{H}_{12}$ ), how many grams of carbon dioxide ( $\text{CO}_2$ ) will be produced?

$$9.75 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{31.998 \text{ g O}_2} = 0.3047 \text{ mol O}_2$$

$$0.3047 \text{ mol O}_2 \times \frac{5 \text{ mol CO}_2}{8 \text{ mol O}_2} = 0.1904 \text{ mol CO}_2$$

$$0.1904 \text{ mol CO}_2 \times \frac{44.009 \text{ g CO}_2}{1 \text{ mol CO}_2} = 8.38 \text{ g CO}_2$$

3. Based on your answers to question #1 and 2, if 5.25 grams of pentane are reacted with 9.75 grams of oxygen, what is the limiting reactant and how many grams of  $\text{CO}_2$  actually form?

**Limiting reactant:  $\text{O}_2$**

**8.38 g  $\text{CO}_2$  actually form**

# Limiting And Excess Reactants Answer Key Extension Questions

**JG Myers**



**Limiting And Excess Reactants Answer Key Extension Questions:**

*Limiting Reactant and Percent Yield*, 2019 This lesson plan covers analyzing chemical reactions in order to determine limiting reactants and excess reactants including calculating the amount of excess reactant calculate the theoretical yield of a reaction and calculating the percent yield of a reaction

## **Limiting And Excess Reactants Answer Key Extrnsion Questions** Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has be much more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Limiting And Excess Reactants Answer Key Extrnsion Questions**," published by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound affect our existence. Throughout this critique, we will delve in to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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