

Limiting & Excess Reactants

Cheney, G., Thomas, T., McClelland, M., Salovey, P., & Mather, T. (2002). Emotions in children's relationships with the classroom teacher.

1. *Journal of the American Medical Association*, 2000; 283: 2686-2692.

- 1.1. **Redox reactions:** $\text{Fe}^{2+} + \text{H}_2\text{O}_2 + 2\text{H}^+ \rightarrow \text{Fe}^{3+} + 2\text{H}_2\text{O}$
In aqueous solution, Fe^{2+} is oxidized to Fe^{3+} by H_2O_2 .
Overall equation:
$$\text{Fe}^{2+} + \text{H}_2\text{O}_2 + 2\text{H}^+ \rightarrow \text{Fe}^{3+} + 2\text{H}_2\text{O}$$

Half-reactions:
Oxidation: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + e^-$
Reduction: $\text{H}_2\text{O}_2 + 2\text{H}^+ + 2e^- \rightarrow 2\text{H}_2\text{O}$
- 1.2. **Reaction of Fe^{2+} with KMnO_4 in acidic solution:**
$$5\text{Fe}^{2+} + \text{MnO}_4^- + 8\text{H}^+ \rightarrow 5\text{Fe}^{3+} + \text{Mn}^{2+} + 4\text{H}_2\text{O}$$

Half-reactions:
Oxidation: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + e^-$
Reduction: $\text{MnO}_4^- + 8\text{H}^+ + 5e^- \rightarrow \text{Mn}^{2+} + 4\text{H}_2\text{O}$
- 1.3. **Reaction of Fe^{2+} with $\text{Cr}_2\text{O}_7^{2-}$ in acidic solution:**
$$6\text{Fe}^{2+} + \text{Cr}_2\text{O}_7^{2-} + 14\text{H}^+ \rightarrow 6\text{Fe}^{3+} + 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$$

Half-reactions:
Oxidation: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + e^-$
Reduction: $\text{Cr}_2\text{O}_7^{2-} + 14\text{H}^+ + 6e^- \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$
- 1.4. **Reaction of Fe^{2+} with IO_3^- in acidic solution:**
$$5\text{Fe}^{2+} + \text{IO}_3^- + 6\text{H}^+ \rightarrow 5\text{Fe}^{3+} + \text{I}^- + 3\text{H}_2\text{O}$$

Half-reactions:
Oxidation: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + e^-$
Reduction: $\text{IO}_3^- + 6\text{H}^+ + 5e^- \rightarrow \text{I}^- + 3\text{H}_2\text{O}$
- 1.5. **Reaction of Fe^{2+} with Br_2 in acidic solution:**
$$2\text{Fe}^{2+} + \text{Br}_2 \rightarrow 2\text{Fe}^{3+} + 2\text{Br}^-$$

Half-reactions:
Oxidation: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + e^-$
Reduction: $\text{Br}_2 + 2e^- \rightarrow 2\text{Br}^-$
- 1.6. **Reaction of Fe^{2+} with H_2O_2 in basic solution:**
$$\text{Fe}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{Fe}^{3+} + 2\text{OH}^-$$

Half-reactions:
Oxidation: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + e^-$
Reduction: $\text{H}_2\text{O}_2 + 2e^- \rightarrow 2\text{OH}^-$
- 1.7. **Reaction of Fe^{2+} with HNO_3 in acidic solution:**
$$3\text{Fe}^{2+} + \text{HNO}_3 + 4\text{H}^+ \rightarrow 3\text{Fe}^{3+} + \text{NO} + 2\text{H}_2\text{O}$$

Half-reactions:
Oxidation: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + e^-$
Reduction: $\text{HNO}_3 + 4\text{H}^+ + 3e^- \rightarrow \text{NO} + 2\text{H}_2\text{O}$
- 1.8. **Reaction of Fe^{2+} with H_2SO_4 in acidic solution:**
$$\text{Fe}^{2+} + \text{H}_2\text{SO}_4 \rightarrow \text{Fe}^{3+} + \text{H}_2\text{S}$$

Half-reactions:
Oxidation: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + e^-$
Reduction: $\text{H}_2\text{SO}_4 + 2e^- + 2\text{H}^+ \rightarrow \text{H}_2\text{S} + 2\text{H}_2\text{O}$
- 1.9. **Reaction of Fe^{2+} with HClO_4 in acidic solution:**
$$\text{Fe}^{2+} + \text{HClO}_4 \rightarrow \text{Fe}^{3+} + \text{HCl}$$

Half-reactions:
Oxidation: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + e^-$
Reduction: $\text{HClO}_4 + 2e^- + 2\text{H}^+ \rightarrow \text{HCl} + \text{H}_2\text{O}$
- 2.0. **Reaction of Fe^{2+} with H_2O_2 in basic solution:**
$$\text{Fe}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{Fe}^{3+} + 2\text{OH}^-$$

Half-reactions:
Oxidation: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + e^-$
Reduction: $\text{H}_2\text{O}_2 + 2e^- \rightarrow 2\text{OH}^-$

Limiting And Excess Reactants Extension Questions Pogil

Xiaolong Qi



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