

# Limiting Reagent Worksheet

## Answers

- ① Given the following reaction:



Balance the reaction.

- ① Suppose 14.8 g of  $\text{C}_3\text{H}_8$  reacts with 3.44 g of  $\text{O}_2$ . Determine the limiting reagent.

$$14.8 \text{ g C}_3\text{H}_8 \times \frac{1 \text{ mol C}_3\text{H}_8}{44.09 \text{ g C}_3\text{H}_8} \times \frac{4 \text{ mol H}_2\text{O}}{1 \text{ mol C}_3\text{H}_8} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 24.2 \text{ g H}_2\text{O}$$

$$3.44 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \times \frac{4 \text{ mol H}_2\text{O}}{5 \text{ mol O}_2} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 1.55 \text{ g H}_2\text{O}$$

Since  $\text{O}_2$  produces less quantity of  $\text{H}_2\text{O}$  than  $\text{C}_3\text{H}_8$  does, it is the limiting reagent.

- ② Determine the amount of  $\text{CO}_2$  produced.

$$3.44 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \times \frac{3 \text{ mol CO}_2}{5 \text{ mol O}_2} \times \frac{44.01 \text{ g CO}_2}{1 \text{ mol H}_2\text{O}} = 2.84 \text{ g CO}_2$$

- ② Given the following reaction:



Balance the reaction.

- ① What is the limiting reagent if 2.2 g of Mg reacts with 4.5 L of  $\text{O}_2$  at STP?

$$2.2 \text{ g Mg} \times \frac{1 \text{ mol Mg}}{24.3 \text{ g Mg}} \times \frac{2 \text{ mol MgO}}{2 \text{ mol Mg}} = 0.0905 \text{ mol MgO}$$

$$4.5 \text{ L O}_2 \times \frac{1 \text{ mol O}_2}{22.4 \text{ L O}_2} \times \frac{2 \text{ mol MgO}}{1 \text{ mol O}_2} = 0.401 \text{ mol MgO}$$

Since Mg produces less MgO than  $\text{O}_2$ , Mg is the limiting reagent.

- ② Determine the amount of MgO produced in grams.

$$0.095 \text{ mol MgO} \times \frac{40.31 \text{ g MgO}}{1 \text{ mol MgO}} = 3.83 \text{ g MgO}$$

# Limiting Reagent Answers Chemfiesta

**Michael Hoelscher**



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## **Table of Contents Limiting Reagent Answers Chemfiesta**

1. Understanding the eBook Limiting Reagent Answers Chemfiesta
  - The Rise of Digital Reading Limiting Reagent Answers Chemfiesta
  - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reagent Answers Chemfiesta
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Limiting Reagent Answers Chemfiesta
  - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reagent Answers Chemfiesta
  - Personalized Recommendations
  - Limiting Reagent Answers Chemfiesta User Reviews and Ratings
  - Limiting Reagent Answers Chemfiesta and Bestseller Lists
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  - Limiting Reagent Answers Chemfiesta Enhanced eBook Features
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  - Adjustable Fonts and Text Sizes of Limiting Reagent Answers Chemfiesta
  - Highlighting and Note-Taking Limiting Reagent Answers Chemfiesta
  - Interactive Elements Limiting Reagent Answers Chemfiesta
- 8. Staying Engaged with Limiting Reagent Answers Chemfiesta
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Limiting Reagent Answers Chemfiesta
- 9. Balancing eBooks and Physical Books Limiting Reagent Answers Chemfiesta
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Limiting Reagent Answers Chemfiesta
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Limiting Reagent Answers Chemfiesta
  - Setting Reading Goals Limiting Reagent Answers Chemfiesta
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Limiting Reagent Answers Chemfiesta
  - Fact-Checking eBook Content of Limiting Reagent Answers Chemfiesta
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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