

SYLLABUS

- (i) Types of chemical changes.
 • *Direct combination* • *decomposition* • *displacement*; • *double decomposition*
(The above to be taught with suitable chemical equations as examples).
- (ii) Energy changes in a chemical change.
Exothermic and endothermic reactions with examples – evolution/absorption of heat, light and electricity.

2.1 CHEMICAL REACTION

A chemical reaction is the process of breaking the *chemical bonds* of the reacting substances (reactants) and making new bonds to form new substances (products).

A chemical bond is the force that holds the atoms of a molecule together, as in a compound.

A chemical change or chemical reaction occurs when particles collide. Collisions occur when reactants are in close contact or by supply of energy. Thus, one or more of the following conditions are necessary for a chemical change :

(i) Mixing (close contact) : In some cases, a chemical reaction occurs when two substances are mixed in their solid states.

Example 1 : Iodine and phosphorus react explosively when brought into close contact.

2 : Lead nitrate (white) and potassium iodide (white) react to make lead iodide (yellow).



(ii) Solution : In some cases, a chemical reaction occurs when substances are mixed in, i.e. molten or aqueous state.

Example 1 : Oxalic acid crystals and sodium carbonate react in water solution only.

2 : Sodium chloride and silver nitrate also react in a solution state to form the precipitate of silver chloride and sodium nitrate.



(iii) Heat : Some chemical reactions occur only on heating.

Example 1 : Copper carbonate decomposes on heating (Δ symbol for heating) into copper oxide and carbon dioxide.



2 : Lead nitrate decomposes on heating leaving yellow residue lead monoxide, brown gas nitrogen dioxide and colourless gas oxygen.



(iv) Light : Some chemical reactions take place by the action of light. These are called *photochemical reactions* or *photolysis*. Molecules of the reactants absorb light energy to get activated, and then react rapidly.

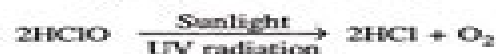
Example 1 : Plants form glucose from carbon dioxide and water in the presence of light.



2 : Hydrogen and chlorine react in the presence of sunlight.



3 : If chlorine water is exposed to sunlight, oxygen is evolved.



Natural Approach To Chemistry Chapter 7 Answers

Lewis Guy Rohrbaugh



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Customize Your Diet to Your Own Unique Body Chemistry For hereditary reasons your metabolism is unique Cutting edge research shows that no single diet works well for everyone the very same foods that keep your best friend slim may keep you overweight and feeling unhealthy and fatigued Now William Wolcott a pioneer in the field of metabolic research has developed a revolutionary weight loss program that allows you to identify your metabolic type and create a diet that suits your individual nutritional needs In The Metabolic Typing Diet Wolcott and acclaimed science writer Trish Fahey provide simple self tests that you can use to discover your own metabolic type and determine what kind of diet will work best for you It might be a low fat high carbohydrate diet filled with pasta and grains or a high fat high protein diet focused on meat and seafood or anything in between By detailing exactly which foods and food combinations are right for you The Metabolic Typing Diet at last reveals the secret to shedding unwanted pounds and achieving optimum vitality with lasting results The Metabolic Typing Diet will enable you to Achieve and maintain your ideal weight Eliminate sugar cravings Enjoy sustained energy and endurance Conquer indigestion fatigue and allergies Bolster your immune system Overcome anxiety depression and mood swings Publications United States. National Bureau of Standards, 1973 **Publications of the National Bureau of Standards** United States. National Bureau of Standards, 1972 **Publications of the National Institute of Standards and Technology ... Catalog** National Institute of Standards and Technology (U.S.), 1971 **Lavoisier** Jean-Pierre Poirier, 1998 Translated from the French and revised and expanded by the author the book provides a rich and detailed account of all facets of Lavoisier s extraordinary career **Catalog of National Bureau of Standards Publications, 1966-1976: pt. 1 Citations and abstracts. v. 2. Key word index (A through L)** United States. National Bureau of Standards. Technical Information and Publications Division, 1978 **Publications of the National Bureau of Standards, 1972 Catalog** United States. National Bureau of Standards, 1973 *Chemical News and Journal of Physical Science*, 1900 *Nature London*, 1871 **Stereochemistry and the Chemistry of Natural Products** Ivor Lionel Finar, 1968 *Nature*, 1892 *Chemical News and Journal of Physical Science* William Crookes, James H. Gardiner, Gerald Druce, H. W. Blood-Ryan, 1905 **Science Education**, 1928 **General Science Quarterly**, 1928 **Paper Makers Monthly Journal**, 1916

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