



Limiting Reactant Identification

6.7 moles of iron (Fe) react with 8.4 moles of oxygen (O_2) according to the balanced equation above. What is the limiting reactant?

Iron (Fe) gmm:

$$(1)(55.9) = 55.9 \text{ g/mol}$$

Oxygen (O_2) gmm:

$$(2)(16.0) = 32.0 \text{ g/mol}$$

mole ratio

$$\frac{6.7 \text{ mol Fe} \times 2 \text{ mol Fe}_2\text{O}_3}{4 \text{ mol Fe}} = 3.35 \text{ mol Fe}_2\text{O}_3$$

Limiting

Exces

$$\frac{8.4 \text{ mol O}_2 \times 2 \text{ mol Fe}_2\text{O}_3}{3 \text{ mol O}_2} = 5.6 \text{ mol Fe}_2\text{O}_3$$

Limiting Reactant Problems And Answers High School

K Morrison



Limiting Reactant Problems And Answers High School:

Limiting Reactant Problems And Answers High School Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has are more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Limiting Reactant Problems And Answers High School**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve in to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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Limiting Reactant Problems And Answers High School Introduction

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