

Limiting Reactant Practice

5 pages + 5-page Answer KEY + Periodic Table

chemistry

LIMITING REACTANTS

1 The reaction below produces ammonia for use in fertilizers.
 $N_2 + 3H_2 \rightarrow 2NH_3$
 If you start with 10 moles of N_2 and 30 moles of H_2 , determine the limiting reactant.

2 Calculate the grams of ammonia produced.

3 How many grams of the excess reactant is produced in a left over?
 To multiply the moles of the limiting reactant by the mole ratio of the excess reactant.

4 Followed moles of excess \times molar mass of excess

5 The reaction below produces carbon dioxide and water from glucose (sugar) and oxygen.
 $C_6H_{12}O_6 + 6O_2 \rightarrow 6CO_2 + 6H_2O$
 If you have 100 grams of glucose and 60 grams of oxygen, determine the limiting reactant.

6 How many grams of water produced?



7 Magnesium chloride is used to create salts for de-icing roads. The reaction with very impure iron ore is produced.
 $Mg + 2FeCl_2 \rightarrow MgCl_2 + 2Fe$
 If you start with 10 moles of magnesium and 2 moles of iron(II) chloride, determine the limiting reactant.

8 Calculate the grams of magnesium chloride produced.

9 How many grams of the excess reactant is produced in a left over?

10 Magnesium chloride is used to create salts for de-icing roads. The reaction with very impure iron ore is produced.
 $Mg + 2FeCl_2 \rightarrow MgCl_2 + 2Fe$
 If you have 100 grams of magnesium and 60 grams of iron(II) chloride, determine the limiting reactant.

11 How much carbon dioxide is produced?

12 The reaction below produces carbon dioxide and water from glucose (sugar) and oxygen.
 $C_6H_{12}O_6 + 6O_2 \rightarrow 6CO_2 + 6H_2O$
 If you have 100 grams of glucose and 60 grams of oxygen, determine the limiting reactant.

13 Calculate the grams of carbon dioxide produced.
 $100 \text{ g } C_6H_{12}O_6 \times \frac{1 \text{ mol } C_6H_{12}O_6}{180 \text{ g } C_6H_{12}O_6} \times \frac{6 \text{ mol } CO_2}{1 \text{ mol } C_6H_{12}O_6} \times \frac{44 \text{ g } CO_2}{1 \text{ mol } CO_2} = 155.56 \text{ g } CO_2$
 $60 \text{ g } O_2 \times \frac{1 \text{ mol } O_2}{32 \text{ g } O_2} \times \frac{6 \text{ mol } CO_2}{6 \text{ mol } O_2} \times \frac{44 \text{ g } CO_2}{1 \text{ mol } CO_2} = 165 \text{ g } CO_2$
 Since 155.56 g < 165 g, glucose is the limiting reactant.

14 How many grams of water produced?
 $100 \text{ g } C_6H_{12}O_6 \times \frac{1 \text{ mol } C_6H_{12}O_6}{180 \text{ g } C_6H_{12}O_6} \times \frac{6 \text{ mol } H_2O}{1 \text{ mol } C_6H_{12}O_6} \times \frac{18 \text{ g } H_2O}{1 \text{ mol } H_2O} = 111.11 \text{ g } H_2O$
 $60 \text{ g } O_2 \times \frac{1 \text{ mol } O_2}{32 \text{ g } O_2} \times \frac{6 \text{ mol } H_2O}{6 \text{ mol } O_2} \times \frac{18 \text{ g } H_2O}{1 \text{ mol } H_2O} = 165 \text{ g } H_2O$
 Since 111.11 g < 165 g, glucose is the limiting reactant.

15 How many moles of carbon dioxide is produced?

16 Calculate the grams of carbon dioxide produced.
 $100 \text{ g } C_6H_{12}O_6 \times \frac{1 \text{ mol } C_6H_{12}O_6}{180 \text{ g } C_6H_{12}O_6} \times \frac{6 \text{ mol } CO_2}{1 \text{ mol } C_6H_{12}O_6} \times \frac{44 \text{ g } CO_2}{1 \text{ mol } CO_2} = 155.56 \text{ g } CO_2$
 $60 \text{ g } O_2 \times \frac{1 \text{ mol } O_2}{32 \text{ g } O_2} \times \frac{6 \text{ mol } CO_2}{6 \text{ mol } O_2} \times \frac{44 \text{ g } CO_2}{1 \text{ mol } CO_2} = 165 \text{ g } CO_2$
 Since 155.56 g < 165 g, glucose is the limiting reactant.

17 How many grams of the excess reactant is produced in a left over?
 To multiply the moles of the limiting reactant by the mole ratio of the excess reactant.

18 Calculate the grams of magnesium chloride produced.
 $100 \text{ g } Mg \times \frac{1 \text{ mol } Mg}{24 \text{ g } Mg} \times \frac{1 \text{ mol } MgCl_2}{1 \text{ mol } Mg} \times \frac{95 \text{ g } MgCl_2}{1 \text{ mol } MgCl_2} = 416.67 \text{ g } MgCl_2$
 $60 \text{ g } FeCl_2 \times \frac{1 \text{ mol } FeCl_2}{152 \text{ g } FeCl_2} \times \frac{1 \text{ mol } MgCl_2}{1 \text{ mol } FeCl_2} \times \frac{95 \text{ g } MgCl_2}{1 \text{ mol } MgCl_2} = 37.5 \text{ g } MgCl_2$
 Since 37.5 g < 416.67 g, iron(II) chloride is the limiting reactant.

19 How much carbon dioxide is produced?

ANSWER KEY INCLUDED

Limiting Reactant Answer Key

CO Houle



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