

# Stoichiometry Limiting Reactant

## Answers

- (1) How many grams of calcium phosphate can be produced from the reaction between 2.50 L of 0.250 M calcium chloride with an excess of phosphoric acid?



$$2.50 \text{ L CaCl}_2 \times \frac{0.250 \text{ mol CaCl}_2}{1 \text{ L CaCl}_2} \times \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{3 \text{ mol CaCl}_2} \times \frac{310 \text{ g Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2} = 64.6 \text{ g Ca}_3(\text{PO}_4)_2$$

- (2) How many milliliters of 1.50 M nitric acid is required to react with 100.0 g of cuprous oxide?



$$100 \text{ g Cu}_2\text{O} \times \frac{1 \text{ mol Cu}_2\text{O}}{143 \text{ g Cu}_2\text{O}} \times \frac{14 \text{ mol HNO}_3}{3 \text{ mol Cu}_2\text{O}} \times \frac{1000 \text{ ml HNO}_3}{1.5 \text{ mol HNO}_3} = 2180 \text{ ml HNO}_3$$

- (3) Consider the following unbalanced reaction:



Balance the reaction. 3.25 g of  $\text{NH}_3$  is allowed to react with 3.5 g of  $\text{O}_2$ .

- (a) Which reactant is the limiting reagent?

$$3.25 \text{ g NH}_3 \times \frac{1 \text{ mol NH}_3}{17.03 \text{ g NH}_3} \times \frac{4 \text{ mol NO}}{4 \text{ mol NH}_3} \times \frac{30.01 \text{ g NO}}{1 \text{ mol NO}} = 5.73 \text{ g NO}$$

$$3.5 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \times \frac{4 \text{ mol NO}}{5 \text{ mol O}_2} \times \frac{30.01 \text{ g NO}}{1 \text{ mol NO}} = 2.63 \text{ g NO}$$

$\text{O}_2$  is the limiting reagent.

- (b) How many grams of NO are formed?

2.63 g of NO are formed.

- (c) How much excess reagent remains after the reaction?

$$3.5 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \times \frac{4 \text{ mol NH}_3}{5 \text{ mol O}_2} \times \frac{17.03 \text{ g NH}_3}{1 \text{ mol NH}_3} = 1.49 \text{ g NH}_3 \text{ used}$$

$$3.25 \text{ g NH}_3 - 1.49 \text{ g NH}_3 = 1.76 \text{ g NH}_3 \text{ remains.}$$

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## **Table of Contents Limiting Reactant Worksheets With Answers Pg 66**

1. Understanding the eBook Limiting Reactant Worksheets With Answers Pg 66
  - The Rise of Digital Reading Limiting Reactant Worksheets With Answers Pg 66
  - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reactant Worksheets With Answers Pg 66
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Limiting Reactant Worksheets With Answers Pg 66
  - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reactant Worksheets With Answers Pg 66
  - Personalized Recommendations
  - Limiting Reactant Worksheets With Answers Pg 66 User Reviews and Ratings

- Limiting Reactant Worksheets With Answers Pg 66 and Bestseller Lists
- 5. Accessing Limiting Reactant Worksheets With Answers Pg 66 Free and Paid eBooks
  - Limiting Reactant Worksheets With Answers Pg 66 Public Domain eBooks
  - Limiting Reactant Worksheets With Answers Pg 66 eBook Subscription Services
  - Limiting Reactant Worksheets With Answers Pg 66 Budget-Friendly Options
- 6. Navigating Limiting Reactant Worksheets With Answers Pg 66 eBook Formats
  - ePub, PDF, MOBI, and More
  - Limiting Reactant Worksheets With Answers Pg 66 Compatibility with Devices
  - Limiting Reactant Worksheets With Answers Pg 66 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Limiting Reactant Worksheets With Answers Pg 66
  - Highlighting and Note-Taking Limiting Reactant Worksheets With Answers Pg 66
  - Interactive Elements Limiting Reactant Worksheets With Answers Pg 66
- 8. Staying Engaged with Limiting Reactant Worksheets With Answers Pg 66
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Limiting Reactant Worksheets With Answers Pg 66
- 9. Balancing eBooks and Physical Books Limiting Reactant Worksheets With Answers Pg 66
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Limiting Reactant Worksheets With Answers Pg 66
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Limiting Reactant Worksheets With Answers Pg 66
  - Setting Reading Goals Limiting Reactant Worksheets With Answers Pg 66
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Limiting Reactant Worksheets With Answers Pg 66
  - Fact-Checking eBook Content of Limiting Reactant Worksheets With Answers Pg 66
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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