Calculating limiting reactants.

Let's look at an example:

$$N_2 + 3H_2 \rightarrow 2NH_3$$



In this reaction, 56 g nitrogen gas (N_2) was added to a chamber containing 14 g hydrogen gas (H_2) . Which was the limiting reactant? $(A_r: H=1, N=14)$

First steps?

Limiting Reactants Study Guide

Wilbert Hutton

Limiting Reactants Study Guide:

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