

Limiting Reactant, Theoretical Yield, Actual Yield, & Percent Yield

$22.99 + 35.45 = 58.44$

24



- If we have 53.2 g of Na and 65.8 g of Cl_2 , what are the limiting reactant and theoretical yield?

$$53.2\text{ g Na} \times \frac{1\text{ mol Na}}{22.99\text{ g Na}} \times \frac{2\text{ mol NaCl}}{2\text{ mol Na}} \times \frac{58.44\text{ g NaCl}}{1\text{ mol NaCl}} = 135.23$$

$$65.8\text{ g Cl}_2 \times \frac{1\text{ mol Cl}_2}{70.9\text{ g Cl}_2} \times \frac{2\text{ mol NaCl}}{1\text{ mol Cl}_2} \times \frac{58.44\text{ g NaCl}}{1\text{ mol NaCl}} = 110.5$$

Limiting Reagent And Percent Yield 12 3

DP Hallahan



Limiting Reagent And Percent Yield 12 3:

Adopting the Beat of Phrase: An Mental Symphony within **Limiting Reagent And Percent Yield 12 3**

In some sort of taken by screens and the ceaseless chatter of instantaneous conversation, the melodic splendor and mental symphony produced by the prepared word often fade into the back ground, eclipsed by the relentless sound and interruptions that permeate our lives. However, nestled within the pages of **Limiting Reagent And Percent Yield 12 3** an enchanting fictional value overflowing with organic thoughts, lies an immersive symphony waiting to be embraced. Constructed by a wonderful composer of language, that fascinating masterpiece conducts viewers on an emotional journey, skillfully unraveling the concealed melodies and profound affect resonating within each cautiously crafted phrase. Within the depths of this touching review, we shall discover the book is central harmonies, analyze their enthralling publishing model, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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