

## CHAPTER

## 02

## SOLUTION

- Solution** : A homogeneous mixture of two or more substances is known as solution.
- Solute** : The substance present in smaller amount in a solution is called solute.
- Solvent** : The substance present in larger amount in a solution is called solvent.
- Binary solution** : These solutions which contain two components are called binary solutions, e.g., salt solution, benzene and toluene.
- Ternary solutions** : Those solutions which contain three components are called ternary solutions, e.g., ethanol + water + acetic acid.
- Aqueous solution** : When solute is dissolved in water, it is called aqueous solution, e.g., sugar solution, ethanol in water.
- Non-aqueous solution** : When solute is dissolved in solvent other than water, it is called non-aqueous solution, e.g., iodine dissolved in alcohol (Tincture of iodine).
- Weight Fraction =  $\frac{\text{weight of solute}}{\text{weight of solution}}$
- Weight percentage  

$$= \frac{\text{weight of solute}}{\text{weight of solution}} \times 100$$

Weight of solute present in 100 grams of solution is called weight percentage

Volume of solute present in 100 ml of the solution is called volume percentage
- Volume percentage  

$$= \frac{\text{Volume of solute}}{\text{Volume of solution}} \times 100$$
- Volume Fraction =  $\frac{\text{Volume of solute}}{\text{Volume of solution}}$
- Molarity (M)  

$$= \frac{\text{No of moles of solute}}{\text{Volume of solution in liters}}$$

(or)  $M = \frac{W}{GMW} \times \frac{1}{V(lit)}$
- Molarity =  $\frac{\text{No of moles of solute}}{\text{Volume of solution in ml}} \times 1000$
- Molarity =  $\frac{\text{Weight of solute}}{\text{GMWt of solute}} \times \frac{1000}{V \text{ in ml}}$   
 (or)  $M = \frac{W}{GMW} \times \frac{1000}{V(ml)}$
- Weight of solute =  $\frac{M \times GMW \times V \text{ in ml}}{1000}$
- Molarity for dilute solutions :  $M_1V_1 = M_2V_2$
- For neutralisation reactions :  $\frac{M_1V_1}{n_1} = \frac{M_2V_2}{n_2}$
- For mixed solution molarity  

$$= \frac{M_1V_1 + M_2V_2 + \dots}{V_1 + V_2 + \dots}$$

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