

## The K<sub>sp</sub> of Lead Iodide, PbI<sub>2</sub>

C12-4-13 & C12-1-01

**Fact:** Many barium compounds are extremely poisonous, yet barium sulfate is routinely swallowed by patients in order to get clearer x-ray pictures of their lower intestines. Why doesn't barium sulfate poison these patients?

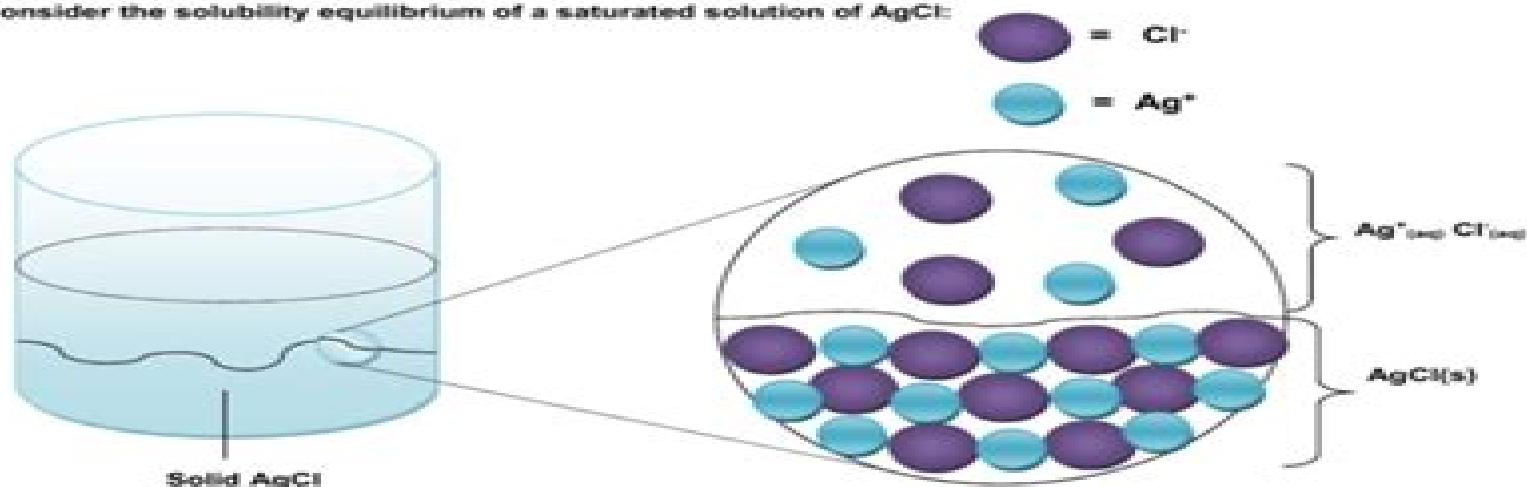
**Fact:** It is possible to saturate a swimming pool with silver bromide and kill off all the nasty little micro-organisms such as algae and bacteria, so that the pool water remains clean and sparkling, yet the water is as pure as drinking water and harmless to people. Since silver ion is a heavy metal ion and hence dangerous to our health in large doses, how is it possible that water saturated with silver bromide doesn't harm us?

The answers to the puzzling questions posed above lies in the low solubility of many salt compounds.

### Introduction:

When ionic solids dissolve, they dissociate into their positive and negative ions that make up the solids. These ions become hydrates and have the same relative proportions when in solution and when solid. The more the solid dissolves, the more the ion concentration increases. This increase in ion concentration allows for the reverse reaction (crystallization) to occur. Eventually the rate of dissolving will equal the rate of crystallization. This state of saturation is called equilibrium. This state can be recognized by a constant color or constant mass.

Consider the solubility equilibrium of a saturated solution of AgCl:



# Ksp Of Lead Iodide Lab Calculations

**SA Adler**



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