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reaction kinetic energy is the heat that is either released or absorbed in the reaction

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web a the solid state has the least amount of entropy associated with it b the gaseous state has the least amount of entropy  
associated with it c the entropy of a system decreases as the matter in the system increases d the entropy of a system  
decreases as the temperature in the system increases e

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disorder that exists in any system when a reaction is exothermic energy will be thought of as a in the reaction if the reaction  
is endothermic energy will be considered a product reactant enthalpy

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when a reaction is endothermic energy will be thought of as a reactant potential energy energy stored in chemical bonds of a

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web 1 kj 1000 j how many j joules are in 1 kj kilo joules positive negative h is for endothermic reactions and for exothermic reactions enthalpy is a state function and is therefore independent of path state hess s law enthalpy of formation hf the h of a formation reaction h hf products hf reactants

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