

EQUILIBRIUM TEST

MULTIPLE CHOICE QUESTIONS

8) Consider the following reaction:



Which one of the following statements is correct?

- A. Since very large, the reaction is irreversible.
- B. Changing the pressure will affect the value of K_c .
- C. It is an very large, the reaction is effectively irreversible.
- D. Changing the temperature will not affect the value of K_c .

9) The following reaction, in which nitrogen trioxide, N_2O_3 , is converted to nitrogen dioxide, NO_2 , occurs at equilibrium:



At a given temperature, the equilibrium constant for this reaction is 0.50. If K_c and the initial concentration of N_2O_3 at equilibrium is 0.500 M, at this temperature, the initial concentrations of NO_2 at equilibrium is

- A. 0.500 M
- B. 1.00 M
- C. 1.50 M
- D. 2.00 M

10) Chloroacetic acid, ClCH_2COOH , reacts with water, H_2O , to form methanoic acid, H_2COOH , H_2COO^- , chloroacetate ions, $\text{ClCH}_2\text{COO}^-$ and hydrochloric acid, HCl . At equilibrium, the HCl and hydrochloric ions, H_3O^+ , according to the equilibrium below:



It is known that increases in K_c of chloroacetic acid is placed in a closed and held overnight and that chloroacetic acid slowly reaches equilibrium. In this scenario, what change is made in the system, which one of the following statements is most likely to result in an increase in the concentrations of H_3O^+ ions?

- A. Change the solution.
- B. Bubble in more CO_2 gas.
- C. Add a few drops of phenolphthalein indicator.
- D. Add a few drops of 0.1 M sodium hydroxide, NaOH , solution.

Multiple Choice Concept Questions On Equilibrium

Jiyuan Zhang



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