

PSI Physics

Name: _____

1. An unknown chemical element is presented by the following formula: A_ZX . What is the name of index Z?
A. Atomic mass number
B. Atomic number
C. Principle quantum number
D. Orbital quantum number
E. Magnetic quantum number
2. An unknown chemical element is presented by the following formula: A_ZX . What is the name of index A?
A. Atomic mass number
B. Atomic number
C. Principle quantum number
D. Orbital quantum number
E. Magnetic quantum number
3. The atomic number is equivalent to which of the following?
A. The number of neutrons in the atom.
B. The number of protons in the atom.
C. The number of nucleons in the atom.
D. The number of α -particles in the atom.
E. None of the above
4. The atomic mass number is equivalent to which of the following?
A. The number of neutrons in the atom.
B. The number of protons in the atom.
C. The number of nucleons in the atom
D. The number of α -particles in the atom.
E. None of the above
5. Which of the following particles has the smallest mass?
A. Proton
B. Electron
C. Neutron
D. Nucleus
E. Nucleon
6. Which of the following statements about the mass of an atom is true?
A. It is evenly divided between the protons and the orbiting electrons.
B. It is evenly divided between the nucleons and the orbiting electrons.
C. It is concentrated in the electron cloud.
D. It is concentrated in the nucleus.
E. It is evenly divided between protons, neutrons and orbiting electrons.
7. Which of the following is correct for the number of neutrons in the nucleus?
A. $N = A - Z$
B. $N = Z - A$
C. $N = Z + A$
D. $N = Z$
E. $N = A$
8. How many electrons are in the ${}^{12}_6C$ atom?
A. 12
B. 6
C. 18
D. 3
E. 9
9. How many nucleons are in the ${}^{20}_{10}Ne$ atom?
A. 12
B. 30
C. 18
D. 10
E. 20
10. How many neutrons are in the ${}^{23}_{11}Na$ atom?
A. 12
B. 11
C. 18
D. 24
E. 9

Multiple Choice Questions Nuclear Physics

Greg Jacobs, Joshua Schulman



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