# Calculating limiting reactants.

Let's look at an example:

$$N_2 + 3H_2 \rightarrow 2NH_3$$



In this reaction, 56 g nitrogen gas  $(N_2)$  was added to a chamber containing 14 g hydrogen gas  $(H_2)$ . Which was the limiting reactant?  $(A_r: H=1, N=14)$ 

First steps?

# **Limiting Reactant 12 3 Answers**

Chris Hren, Peter J. Mikulecky

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