

Limiting Reactants Practice KEY

1. How many grams of aluminum chloride could be produced from 35.5 g of aluminum and 39.0 g of chlorine gas?



$$35.5 \text{ g Al} \times \frac{1 \text{ mol Al}}{26.982 \text{ g Al}} \times \frac{3 \text{ mol Cl}_2}{2 \text{ mol Al}} \times \frac{70.906 \text{ g Cl}_2}{1 \text{ mol Cl}_2} = 140. \text{ g Cl}_2$$

- ∞ This means 35.5 g Al would react with 140. g Cl₂
- ∞ Only 39.0 g Cl₂ were available, so it is the limiting reactant.
- ∞ Use the mass of Cl₂ in the reaction to determine the mass of the AlCl₃ formed.

$$39.0 \text{ g Cl}_2 = \text{ ? } \text{ g AlCl}_3$$

$$39.0 \text{ g Cl}_2 \times \frac{1 \text{ mol Cl}_2}{70.906 \text{ g Cl}_2} \times \frac{2 \text{ mol AlCl}_3}{3 \text{ mol Cl}_2} \times \frac{133.341 \text{ g AlCl}_3}{1 \text{ mol AlCl}_3} =$$

Answer: 48.9 g AlCl₃

2. Identify the limiting reagent when 10.8 g water reacts with 4.7 g Na to produce NaOH and H₂.

To solve, set one reactant to the other and compare what is calculated to the actual amount in the reaction to determine the theoretical limiting reactant.

$$10.8 \text{ g H}_2\text{O} = \text{ ? } \text{ g Na}$$

Or

$$4.7 \text{ g Na} = \text{ ? } \text{ g H}_2\text{O}$$



$$10.8 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.015 \text{ g H}_2\text{O}} \times \frac{2 \text{ mol Na}}{2 \text{ mol H}_2\text{O}} \times \frac{22.990 \text{ g Na}}{1 \text{ mol Na}} = 13.8 \text{ g Na}$$

- ∞ This means that 10.8 g will react completely with 13.8 g Na, but only 4.7 g Na was available to this reaction. Therefore, Na is the limiting reactant in this problem.

Limiting Reactant Problems Answer Key

SB Merriam



Limiting Reactant Problems Answer Key:

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Mapping College Chemistry Stephen DeMeo, 2019-03-01 This text is a chemistry problem solving resource appropriate for teachers and their students who are enrolled in high school Advanced Placement Chemistry or in a first year college General Chemistry course The book incorporates a chemistry problem solving plan one that uses an innovative graphic organizer strategy The strategy successfully evaluated with students combines problem solving processes with chemical concepts that will allow students to solve the most common and difficult problems encountered in the first year of chemistry Topical problem solving will focus on limiting reactant stoichiometry identifying types of chemical reactions equilibrium acid base equilibria and electrochemistry Why would this resource be of interest to chemistry students To be successful to get into a well known college medical school physical therapy or graduate program often requires that students get an A in your pre requisite Introductory General Chemistry course To make matters worse many college professors feel that only a few students should get A grades and therefore they give difficult exams that many students fail this is the weeding out process that every pre health student is apprehensive about To succeed in this competitive environment entails not just studying harder or longer it means re organizing textbook content so that it is meaningful to the student This is the first text of its kind to employ a reliable research based strategy that incorporates a decision based visual tool to solve chemistry textbook problems ones that can make or break a career

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fundamental chemical concepts most often covered in an introductory nonmajors course Covering all the standard topics Matthew Johll capitalizes on the surge of interest in the scientific investigation of crime as sparked by CSI and other television shows bringing together the theme of forensic science and the fundamentals of chemistry in ways that are effective and accessible for students This edition features refined explanations of the chemical concepts which are the core of the book as well as a more thoroughly integrated forensic theme updated features and an expanded media supplements package

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Table of Contents Limiting Reactant Problems Answer Key

1. Understanding the eBook Limiting Reactant Problems Answer Key
 - The Rise of Digital Reading Limiting Reactant Problems Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reactant Problems Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting Reactant Problems Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reactant Problems Answer Key
 - Personalized Recommendations
 - Limiting Reactant Problems Answer Key User Reviews and Ratings
 - Limiting Reactant Problems Answer Key and Bestseller Lists
5. Accessing Limiting Reactant Problems Answer Key Free and Paid eBooks
 - Limiting Reactant Problems Answer Key Public Domain eBooks
 - Limiting Reactant Problems Answer Key eBook Subscription Services
 - Limiting Reactant Problems Answer Key Budget-Friendly Options

6. Navigating Limiting Reactant Problems Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Limiting Reactant Problems Answer Key Compatibility with Devices
 - Limiting Reactant Problems Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limiting Reactant Problems Answer Key
 - Highlighting and Note-Taking Limiting Reactant Problems Answer Key
 - Interactive Elements Limiting Reactant Problems Answer Key
8. Staying Engaged with Limiting Reactant Problems Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting Reactant Problems Answer Key
9. Balancing eBooks and Physical Books Limiting Reactant Problems Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reactant Problems Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Limiting Reactant Problems Answer Key
 - Setting Reading Goals Limiting Reactant Problems Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Limiting Reactant Problems Answer Key
 - Fact-Checking eBook Content of Limiting Reactant Problems Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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