

Limiting Reagent & Percent Yield Practice Worksheet

1. When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.

a. Write the balanced equation for the reaction given above:



- b. If 15 grams of copper (II) chloride react with 20 grams of sodium nitrate, how much sodium chloride can be formed?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.362 \text{ g CuCl}_2} \times \frac{2 \text{ mol NaCl}}{1 \text{ mol CuCl}_2} \times \frac{58.443 \text{ g NaCl}}{1 \text{ mol NaCl}} = 13 \text{ g NaCl}$$

$$20 \text{ g NaNO}_3 \times \frac{1 \text{ mol NaNO}_3}{84.997 \text{ g NaNO}_3} \times \frac{2 \text{ mol NaCl}}{2 \text{ mol NaNO}_3} \times \frac{58.443 \text{ g NaCl}}{1 \text{ mol NaCl}} = 13.8 \text{ g NaCl}$$

Only 13 g can be formed

- c. What is the limiting reagent for the reaction in #2? **copper (II) chloride**

- d. How many grams of copper (II) nitrate is formed?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.452 \text{ g CuCl}_2} \times \frac{1 \text{ mol Cu(NO}_3)_2}{1 \text{ mol CuCl}_2} \times \frac{187.467 \text{ g Cu(NO}_3)_2}{1 \text{ mol Cu(NO}_3)_2} = 20.9 \text{ g Cu(NO}_3)_2$$

- e. How much of the excess reagent is left over in this reaction?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.452 \text{ g CuCl}_2} \times \frac{2 \text{ mol NaNO}_3}{1 \text{ mol CuCl}_2} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 19 \text{ g NaNO}_3$$

$$20 \text{ g NaNO}_3 - 19 \text{ g NaNO}_3 = 1 \text{ g NaNO}_3$$

- f. If 11.3 grams of sodium chloride are formed in the reaction described in problem #2, what is the percent yield of this reaction?

$$\text{Percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100 \text{ percent}$$

$$11.3/13.0 \times 100\% = 86.9\%$$

2. When lead (II) nitrate reacts with sodium iodide, sodium nitrate and lead (II) iodide are formed.

a. Balance the following equation:



- b. If I start with 25.0 grams of lead (II) nitrate and 15.0 grams of sodium iodide, how many grams of sodium nitrate can be formed?

$$25.0 \text{ g Pb(NO}_3)_2 \times \frac{1 \text{ mol Pb(NO}_3)_2}{331.214 \text{ g Pb(NO}_3)_2} \times \frac{2 \text{ mol NaNO}_3}{1 \text{ mol Pb(NO}_3)_2} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 12.8 \text{ g NaNO}_3$$

$$15.0 \text{ g NaI} \times \frac{1 \text{ mol NaI}}{149.89 \text{ g NaI}} \times \frac{2 \text{ mol NaNO}_3}{2 \text{ mol NaI}} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 8.51 \text{ g NaNO}_3$$

8.51 g NaNO₃ can be made

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Ensheng Dong



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Table of Contents Limiting Reagent And Percent Yield Study Packet

1. Understanding the eBook Limiting Reagent And Percent Yield Study Packet
 - The Rise of Digital Reading Limiting Reagent And Percent Yield Study Packet
 - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reagent And Percent Yield Study Packet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting Reagent And Percent Yield Study Packet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reagent And Percent Yield Study Packet
 - Personalized Recommendations
 - Limiting Reagent And Percent Yield Study Packet User Reviews and Ratings
 - Limiting Reagent And Percent Yield Study Packet and Bestseller Lists
5. Accessing Limiting Reagent And Percent Yield Study Packet Free and Paid eBooks
 - Limiting Reagent And Percent Yield Study Packet Public Domain eBooks
 - Limiting Reagent And Percent Yield Study Packet eBook Subscription Services
 - Limiting Reagent And Percent Yield Study Packet Budget-Friendly Options
6. Navigating Limiting Reagent And Percent Yield Study Packet eBook Formats
 - ePub, PDF, MOBI, and More
 - Limiting Reagent And Percent Yield Study Packet Compatibility with Devices
 - Limiting Reagent And Percent Yield Study Packet Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limiting Reagent And Percent Yield Study Packet
 - Highlighting and Note-Taking Limiting Reagent And Percent Yield Study Packet
 - Interactive Elements Limiting Reagent And Percent Yield Study Packet
8. Staying Engaged with Limiting Reagent And Percent Yield Study Packet

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Limiting Reagent And Percent Yield Study Packet
- 9. Balancing eBooks and Physical Books Limiting Reagent And Percent Yield Study Packet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reagent And Percent Yield Study Packet
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Limiting Reagent And Percent Yield Study Packet
 - Setting Reading Goals Limiting Reagent And Percent Yield Study Packet
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Limiting Reagent And Percent Yield Study Packet
 - Fact-Checking eBook Content of Limiting Reagent And Percent Yield Study Packet
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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