Name:	KEY		Date:	Perio	d:
1	N ₂ + _3_H ₂ → _1_NH ₃	ting React	tant Pract	ice	
	For the above reaction, how mailimiting reactant? What is the ex Ng+3Hg -> 2NHg 29g 25g ?9	cess reactant? How re		ess are left over?	0/2/4/2000 0.000 0.000 0.000
259 Hzx Bucess Hz	2.029 Hz - 259 Hz	17.03910H3 1mol 10H3 134.010910H3	140 59 PH3	Excess Reactant: _ Excess Reactant (g): _	Ha
2.	Mg + 2. HCl → Mg For the above reaction, what vol HCl? What is the limiting reactar	Cl ₂ + H ₂ lume of hydrogen (at 5 st? What is the excess	TP) is produced from	reacting 50.0 g of Ma	and 75 g of
50.09	+ 2HC(-> mgClz 1 73g	? L	g HC1 x I mol F	Imol Mg	-= 2.1 mol H2
1.0	mol H2 x 2.029H2 = 4 mol H2 x 2.029H2 = 2 lmol H2	1 12115		Limiting Reactant:	HCI_
	50gmg (2.02gHz)=26	29 Hz 5 47.	L H2	Hyproduced (g): _ Hz produced	
3AaN02 +	For the above reaction, you reac silver phosphate are produced?	t 200.0 g of silver nitr	ate with 200.0 g of soc	dium phosphate. How	many grams of H ₂
200.0g	200.0g ?g	X 3 mol Ag 1003	418.58gAgs × 1mo1AgsPO4	POY = \Left. Z	Aq N03
200a Nad	.0gAgNO3x 1mol AgNO3 169.9gAgNO PO4x 1mol Na3PO4_x 1 169.94gNa3PO4 1 = 200g Na3PO4 - 200g N	1001 A93P04 419	AgePDY	Excess Reactant: _ Excess Reactant (g): _ AHF. Produced (g):	Na, PDY 135.79 Na, POY 164.259 A9, PO
Bylless :	= 2009 Na3P04 - 2009 N	Day PD4 (164.25	510.79 Agaroy 9 Agaroy) =	135.79 Na	704

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