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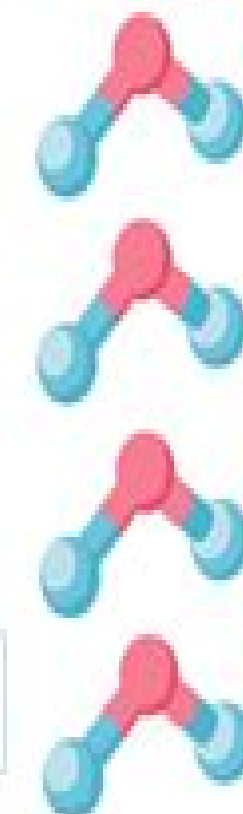
if you begin this reaction with 4 hydrogen molecules and 4 oxygen molecules, hydrogen runs out and limits the number of water molecules that can be produced.



Hydrogen is the limiting reactant.



Oxygen is the excess reactant



There are 2 unreacted oxygen molecules remaining after the reaction.

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Limiting Reactants And Excess Packet

Rob Lewis, Wynne Evans



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Chemistry Rob Lewis, Wynne Evans, 2018-03-31 The fifth edition of this engaging and established textbook provides students with a complete course in chemical literacy and assumes minimal prior experience of science and maths Written in an accessible and succinct style this book offers comprehensive coverage of all the core topics in organic inorganic and physical chemistry Topics covered include bonding moles solutions and solubility energy changes equilibrium organic compounds and spectroscopy Each unit contains in text exercises and revision questions to consolidate learning at every step and is richly illustrated with diagrams and images to aid understanding This popular text is an essential resource for students who are looking for an accessible introductory textbook It is also ideal for non specialists on courses such as general science engineering environmental health or life sciences New to this Edition A foreword by Professor Sir John Meurig Thomas FRS former Director of the Royal Institution Three additional units on Gibbs Energy Changes Organic Mechanisms and Fire and Flame Accompanying online resources for this title can be found at bloomsburyonlineresources.com/chemistry/5e These resources are designed to support teaching and learning when using this textbook and are available at no extra cost

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