Name : _____ Date : _____

# Limiting Reagent Worksheet

Answers

(1) Given the following reaction:

$$____\ C_3H_8 + \underline{\ 5\ }\ O_2 \rightarrow \underline{\ 3\ }\ CO_2 + \underline{\ 4\ }\ H_2O$$

Balance the reaction.

(a) Suppose 14.8 g of $C_3H_8$ reacts with 3.44 g of $O_2$. Determine the limiting reagent.

$$14.8\ g\ C_3H_8 \times \frac{1\ mol\ C_3H_8}{44.09\ g\ C_3H_8} \times \frac{4\ mol\ H_2O}{1\ mol\ C_3H_8} \times \frac{18.02\ g\ H_2O}{1\ mol\ H_2O} = 24.2\ g\ H_2O$$

$$3.44\ g\ O_2 \times \frac{1\ mol\ O_2}{32\ g\ O_2} \times \frac{4\ mol\ H_2O}{5\ mol\ O_2} \times \frac{18.02\ g\ H_2O}{1\ mol\ H_2O} = 1.55\ g\ H_2O$$

Since $O_2$ produces less quantity of $H_2O$ than $C_3H_8$ does, it is the limiting reagent.

(b) Determine the amount of $CO_2$ produced.

$$3.44\ g\ O_2 \times \frac{1\ mol\ O_2}{32\ g\ O_2} \times \frac{3\ mol\ CO_2}{5\ mol\ O_2} \times \frac{44.01\ g\ CO_2}{1\ mol\ H_2O} = 2.84\ g\ CO_2$$

(2) Given the following reaction:

$$\underline{\ 2\ }\ Mg + ____\ O_2 \rightarrow \underline{\ 2\ }\ MgO$$

Balance the reaction.

(a) What is the limiting reagent if 2.2 g of Mg reacts with 4.5 L of $O_2$ at STP?

$$2.2\ g\ Mg \times \frac{1\ mol\ Mg}{24.3\ g\ Mg} \times \frac{2\ mol\ MgO}{2\ mol\ Mg} = 0.0905\ mol\ MgO$$

$$4.5\ L\ O_2 \times \frac{1\ mol\ O_2}{22.4\ L\ O_2} \times \frac{2\ mol\ MgO}{1\ mol\ O_2} \times = 0.401\ mol\ MgO$$

Since Mg produces less MgO than $O_2$, Mg is the limiting reagent.

(b) Determine the amount of MgO produced in grams.

$$0.095\ mol\ MgO \times \frac{40.31\ g\ MgO}{1\ mol\ MgO} = 3.83\ g\ MgO$$

# Limiting Reagent Worksheet 2answer Key

VM Jensen

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