

125. The ribosomes responsible for protein synthesis are present in the cell
a) Floating in the cytosol
b) Localized in the nucleus
c) Bound to rough endoplasmic reticulum
d) Both a and c
126. _____ Enzyme need a primer for the initiation of its function.
a) RNA polymerase
b) DNA polymerase
c) Primase
d) Ligase
127. The following histone proteins form a nucleosome complex except
a) H1
b) H2A
c) H2B
d) H3
128. The bond that is formed between two monosaccharide units is called
a) Ionic bond
b) hydrogen bond
c) peptide bond
d) Glycosidic bond
129. They already _____ some of the old ones and _____ them more comfortable.
a) Repair, make
b) repaired, made
c) repaired, make
d) repair, made
130. I was born in Peshawar but _____ most of my childhood in the Mardan.
a) Spends
b) have spent
c) was spending
d) is spending
131. Which oxides of "K" contain more oxygen than its normal oxide?
a) Per-oxide
b) super oxide
c) both contain equal quantity
d) none of the above
132. A gas decolorizes alkaline KMnO_4 solution, but does not give any PPT with ammoniacal AgNO_3
a) Methane
b) ethylene
c) ethane
d) None of the above
133. Why ethanoic acid is a stronger acid in the liquid ammonia than in water?
a) Ammonia is stronger base than in water
b) Ethanoic acid molecules form H-bonding with water
c) Ethanoic acid is more soluble in liquid ammonia than in water
d) None of the above
134. Which ions are used as catalyst in the reaction between persulphate ions and iodide ions?
a) Lead
b) iron
c) copper
d) chromium
135. Which one is stronger nucleophile?
a) $\text{C}_2\text{H}_5\text{O}^-$
b) $\text{C}_2\text{H}_5\text{S}^-$
c) both are equally strong
d) none of the above
136. Which one of the following elements has the largest second ionization energy?
a) O
b) F
c) Na
d) Ne
137. An α -particle is accelerated through a potential difference of 10^6 volts. Its K.E is
a) 1MeV
b) 2MeV
c) 4MeV
d) 8MeV
138. If there are "n" capacitors each of capacity "C" connected in parallel to "V" volt source then energy stored is equal to
a) CV
b) $\frac{1}{2} nCV^2$
c) CV^2
d) $CV^2/2n$
139. The electric field strength between a pair of plates is "E", if the separation of the plates is doubled and

- potential difference between the plates is increased by factor of four, the new field strength is
a) E
b) 2E
c) 4E
d) 8E
140. Two satellites of masses "3M" and "M" orbit the earth in a circular orbit of radius "r" and "3r" respectively, the ratio of their speed is
a) 1:1
b) $\sqrt{3}:1$
c) 3:1
d) 9:1
141. The optimum pH of enzyme urease is
a) 7.8-8.7
b) 7.0
c) 4.5
d) 8.0
142. Which statement about chlorophyll is not true?
a) It contains terminal carbonyl group
b) It contains phytyl tail
c) It contains porphyrin ring
d) It contains magnesium
143. In humans the disease symptoms develop during the
a) Log phase
b) lag phase
c) growth phase
d) decline phase
144. Independent gametophyte and sporophyte are found in
a) Selaginella
b) Polytrichum
c) Ectocarpus
d) liverworts
145. *Tmesipteris* is an example of
a) Horsetail
b) clubmosses
c) psilopsida
d) Pteropsida
146. The larva formed during the life cycle of Annelida is
a) Glochidium larva
b) Bipinnaria larva
c) trochophore larva
d) tornaria larva
147. Ebner's gland on the dorsal surface of the tongue secrete an enzyme
a) Amylase
b) Ptyalin
c) Lingual lipase
d) both a and b
148. Antibodies consists of _____ polypeptide chains.
a) 2
b) 4
c) 6
d) 8
149. _____ you win first place, you will receive a prize.
a) Whenever
b) if
c) unless
d) so forth
150. The train was _____.
a) Halt
b) halted
c) had halted
d) has halted
151. Which of the following species has the maximum number of unpaired electrons?
a) O_2
b) O^{+2}
c) O^{-2}
d) O_2^{-2}
152. A mixture of 10cm^3 of oxygen and 50cm^3 of hydrogen is sparked continuously. What is the maximum theoretical decrease in volume?
a) 10cm^3
b) 15cm^3
c) 20cm^3
d) 30cm^3
153. The oxidation state of nitrogen in NH_4NO_3 are
a) -3 and +5
b) +5 and -3
c) -3 and -3
d) zero
154. Which equation relates to the first ionization energy of bromine?
a) $\text{Br}_{(g)} \rightarrow \text{Br}_{(g)}^+ + 1e^-$
b) $\text{Br}_{(g)} \rightarrow \text{Br}_{(g)}^- + 1e^-$
c) $\frac{1}{2} \text{Br}_{2(g)} \rightarrow \text{Br}_{(g)} + 1e^-$
d) $\frac{1}{2} \text{Br}_{2(g)} \rightarrow \text{Br}_{(g)}^+ + 1e^-$
155. Co-ordination number of $[\text{Co}(\text{en})_2 \text{Cl}_2]$ is
a) -2
b) 6
c) 4
d) None of the above
156. An olefin "X" on ozonolysis gives $\text{CH}_3\text{CH}_2\text{COCH}_3$ and CH_3COCH_3 . The IUPAC name of X is

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