# Calculations need to be based on the limiting reactant.

 Example 2: What is the mass of sulfur dioxide that can be produced by the reaction of 95.6 g carbon disulfide with 100. g oxygen?

$$CS_2 + 3O_2 \rightarrow CO_2 + 2SO_2$$

$$\frac{95.6g\ CS_2}{76.2g\ CS_2} \times \frac{1mol\ CS_2}{1mol\ CS_2} \times \frac{2mol\ SO_2}{1mol\ CS_2} \times \frac{64.1g\ SO_2}{1mol\ SO_2} = 161g\ SO_2$$

$$\frac{100.g \ O_2}{32.0g \ O_2} \times \frac{1mol \ O_2}{32.0g \ O_2} \times \frac{2mol \ SO_2}{3mol \ O_2} \times \frac{64.1g \ SO_2}{1mol \ SO_2} = 134g \ SO_2$$

134 g of SO<sub>2</sub> can be produced in this reaction.

# **Limiting Reactant Answers**

**KJ Lindholm-Leary** 



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