

# 7. GROUPS 1 & 2

## SECTION B

Answer ALL the questions. Write your answers in the spaces provided.

- 1 Magnesium oxide, known as magnesia, can be produced by the thermal decomposition of magnesium carbonate.

The equation for the reaction is



- \*(a) Other Group 2 carbonates also undergo thermal decomposition.

State and explain the trend in thermal stability of carbonates down Group 2.

(4)

**Thermal stability increases down the group**

**Because ionic radius increases as you go down the group**

**This results in less polarisation (distortion) of the carbonate ion**

**And less weakening of the covalent bonds (C-O) within the carbonate ion**

- \*(c) State why carrying out a flame test is **not** an appropriate way to confirm the presence of magnesium ions in the sample of magnesia.

(1)

**Magnesium ions give no flame colour**

**OR**

**Energy is emitted with Mg ions is outside of the visible region**

(Total for Question 1 = 15 marks)

# March 2013 Past Papers Science Edexcel

**David Spiegelhalter**



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