

CHIMIE

(7 points)

Exercice n°1 (2,75 pts)

Dans un récipient de volume V , on introduit **0,2 mol** d'ammoniac NH_3 gazeux et **0,7 mol** de dioxygène O_2 à une température T_1 .

A l'équilibre chimique, il se forme **0,12 mol** de dioxyde d'azote gazeux NO_2 selon une réaction endothermique d'équation :



- 1*) a- En se servant d'un tableau descriptif d'évolution, déterminer, en mol, la composition finale du mélange.
b- En déduire la valeur du taux d'avancement final τ_{11} de la réaction étudiée.
- 2*) A pression constante, on fait varier la température du mélange obtenu à l'équilibre précédent.
Le taux d'avancement final de la réaction étudiée prend une nouvelle valeur $\tau_{12} = 0,4$.
a- Déterminer le nombre de mole total de gaz dans le nouvel état d'équilibre.
b- Préciser, en justifiant la réponse, si on a augmenté ou diminué la température du milieu réactionnel.
- 3*) Dans quel sens se déplace l'équilibre si on augmente la pression du mélange réactionnel obtenu à la température T_1 ?

Exercice n°2 (4,25 pts)

On donne : Le produit ionique de l'eau : $K_e = 10^{-14}$.

- 1*) a- Définir une base selon la théorie de Brønsted.
b- Etablir, pour un couple acide base AH/A^- , la relation reliant la constante d'acidité K_a , la constante de basicité K_b et le produit ionique de l'eau K_e .
- 2*) a- Compléter (sur la page 3 de la feuille annexe) le tableau suivant en indiquant la formule de la base ou de l'acide conjugué de chaque composé donné et en calculant les valeurs de K_b et du $\text{p}K_a$.

composé acide	HNO_2
Base conjuguée	NH_3	ClO^-	HCOO^-
$\text{p}K_a$	7,5	3,3
K_b	$1,58 \cdot 10^{-5}$	$6,31 \cdot 10^{-11}$

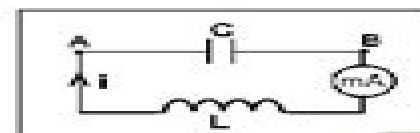
- b- Classer les bases de ces couples acide-base par ordre de force décroissante.
- 3*) L'acide nitreux HNO_2 , de constante d'acidité K_{a1} , réagit avec la base HCOO^- de constante d'acidité K_{a2} .
a- Ecrire l'équation de cette réaction.
b- Montrer que la constante d'équilibre peut s'écrire sous la forme $K = 10^{\text{p}K_{a1} + \text{p}K_{a2} - \text{p}K_e}$.
Calculer sa valeur. Vérifier la comparaison des forces des deux bases effectuée à la question 2-b.
- 4*) Si on part d'un mélange équimolaire contenant $n_0 = 7 \cdot 10^{-2}$ mol de chacune des entités intervenant dans la réaction précédente.
a- Déterminer la composition du mélange à l'équilibre.
b- On agit sans variation de volume considérable sur la quantité d'acide nitreux HNO_2 obtenue à l'équilibre, pour rendre la quantité de HCOO^- égale à $6 \cdot 10^{-2}$ mol.
Préciser, en le justifiant, s'il s'agit d'une augmentation ou d'une diminution de la quantité de HNO_2 .

PHYSIQUE

(13 points)

Exercice n° 1 : (5,5 pts)

Un condensateur de capacité C , préalablement chargé sous une tension $U_{AB} > 0$, est lié à une bobine purement inductive d'inductance $L = 0,5 \text{ H}$ et un milliampèremètre de résistance négligeable.
Des oscillations électriques prennent naissance dans le circuit.



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WJ Hussar



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Table of Contents Js2 Sagetion Math 2014

1. Understanding the eBook Js2 Sagetion Math 2014
 - The Rise of Digital Reading Js2 Sagetion Math 2014
 - Advantages of eBooks Over Traditional Books
2. Identifying Js2 Sagetion Math 2014
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Js2 Sagetion Math 2014
 - User-Friendly Interface
4. Exploring eBook Recommendations from Js2 Sagetion Math 2014
 - Personalized Recommendations
 - Js2 Sagetion Math 2014 User Reviews and Ratings
 - Js2 Sagetion Math 2014 and Bestseller Lists
5. Accessing Js2 Sagetion Math 2014 Free and Paid eBooks
 - Js2 Sagetion Math 2014 Public Domain eBooks
 - Js2 Sagetion Math 2014 eBook Subscription Services
 - Js2 Sagetion Math 2014 Budget-Friendly Options
6. Navigating Js2 Sagetion Math 2014 eBook Formats
 - ePub, PDF, MOBI, and More
 - Js2 Sagetion Math 2014 Compatibility with Devices
 - Js2 Sagetion Math 2014 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Js2 Sagetion Math 2014
 - Highlighting and Note-Taking Js2 Sagetion Math 2014
 - Interactive Elements Js2 Sagetion Math 2014
8. Staying Engaged with Js2 Sagetion Math 2014

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Js2 Sagetion Math 2014
- 9. Balancing eBooks and Physical Books Js2 Sagetion Math 2014
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Js2 Sagetion Math 2014
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Js2 Sagetion Math 2014
 - Setting Reading Goals Js2 Sagetion Math 2014
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Js2 Sagetion Math 2014
 - Fact-Checking eBook Content of Js2 Sagetion Math 2014
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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