

CHEMICAL THERMODYNAMICS

- 1) In open system, system and surroundings exchange.....
 (a) Energy only (b) Matter only
(c) Energy and matter (d) Neither energy nor matter
- 2) An isothermal expansion takes place at.....
 (a) constant pressure (b) constant volume
(c) constant temperature (d) constant temperature & pressure
- 3) In a closed thermos flask few cubes of ice are placed. It is example of.....
 (a) closed system (b) open system
(c) isolated system (d) thermodynamics system
- 4) In an adiabatic compression, the internal energy of the system.....
 (a) decreases **(b) increases** (c) increases or decreases (d) none
- 5) Identify the intensive quantities from the following.....
 (a) volume (b) enthalpy **(c) pressure** (d) none
- 6) Which of the following is extensive property?
 (a) mole (b) mass (c) volume **(d) all are correct**
- 7) The heat change in a chemical reaction at constant pressure is.....
(a) ΔH (b) ΔU (c) ΔT (d) ΔV
- 8) ΔU is always positive when.....
(a) system absorbs heat and work is done on it
 (b) system emits heat and work is done on it
 (c) system emits heat and no work is done on it
 (d) system absorbs heat and work is done by it
- 9) If in a chemical reaction, the products have less energy than the reactants, that reaction is.....
 (a) a reversible reaction **(b) an exothermic reaction**
 (c) an isothermal reaction (d) an endothermic reaction
- 10) A small piece of Na metal is dropped into water. Na bursts into flame, the reaction is:
 (a) neutralisation (b) reversible **(c) exothermic** (d) endothermic
- 11) For the reaction: $C_3H_8(g) + 5O_2(g) \rightarrow 3CO_2(g) + 4H_2O(l)$ at constant temperature, $\Delta H - \Delta U$ is
 (a) $+RT$ **(b) $-3RT$** (c) $+3RT$ (d) $-RT$
- 12) When NH_4Cl is dissolved in water, the solution becomes cold. The change is?
(a) endothermic (b) exothermic (c) super cooling (d) none

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Venkanna,B. V. Swati,2010-07 This book provides an in depth discussion of the principles of thermodynamics It focuses on engineering applications of theory and sound techniques for solving thermodynamic problems The book presents the fundamental concepts of thermodynamics and describes the theory of work and heat The text covers in detail the first law and the second law of thermodynamics with their applications It also explains the concepts of entropy and availability and irreversibility In addition the book presents thermodynamic properties of pure substances ideal gases and mixtures of ideal gases as well as real gases This book is designed for undergraduate students of mechanical engineering industrial and production engineering automobile engineering and aeronautical engineering for their courses in thermodynamics Key Features Presents the text in a simple and elegant manner to enable the students to grasp the essentials of the subject easily and quickly Covers all types of problems of various difficulty levels Includes more than 300 worked out examples and a large number of end of chapter exercises Provides solutions to several model question papers at the end of the book

Thermodynamics Rumi Michael Leigh,2023-05-24 *Thermodynamics Things You Should Know Questions and Answers* explains the principles of thermodynamics in a question and answer format The book introduces energy work heat and the basic laws of thermodynamics It describes the concepts of entropy enthalpy internal energy and equilibrium and shows how these ideas connect to physical and chemical processes Key topics include the first and second laws of thermodynamics the role of entropy in natural systems and the behavior of gases as described by equations of state The text also examines heat transfer efficiency and applications of thermodynamics in engines refrigeration and power systems Worked examples and explanations show how thermodynamic concepts are applied in solving problems and understanding physical phenomena This book will interest students engineering learners and those studying physics who want to understand thermodynamics and its applications *Solutions Manual for Thermodynamics* Jack Philip Holman,1974

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