

## Section 4.5

# Types of Aqueous Solutions and Solubility

## Solubility Rules for Ionic Compounds in Water

TABLE 4.1 Solubility Rules for Ionic Compounds in Water

Compounds Containing the Following Ions Are Generally Soluble	Exceptions
$\text{Li}^+$ , $\text{Na}^+$ , $\text{K}^+$ , and $\text{NH}_4^+$	None
$\text{NO}_3^-$ and $\text{C}_2\text{H}_3\text{O}_2^-$	None
$\text{Cl}^-$ , $\text{Br}^-$ , and $\text{I}^-$	When these ions pair with $\text{Ag}^+$ , $\text{Hg}_2^{2+}$ , or $\text{Pb}^{2+}$ , the resulting compounds are insoluble.
$\text{SO}_4^{2-}$	When $\text{SO}_4^{2-}$ pairs with $\text{Sr}^{2+}$ , $\text{Ba}^{2+}$ , $\text{Pb}^{2+}$ , $\text{Ag}^+$ , or $\text{Ca}^{2+}$ , the resulting compound is insoluble.
Compounds Containing the Following Ions Are Generally Insoluble	Exceptions
$\text{OH}^-$ and $\text{S}^{2-}$	When these ions pair with $\text{Li}^+$ , $\text{Na}^+$ , $\text{K}^+$ , or $\text{NH}_4^+$ , the resulting compounds are soluble.
	When $\text{S}^{2-}$ pairs with $\text{Ca}^{2+}$ , $\text{Sr}^{2+}$ , or $\text{Ba}^{2+}$ , the resulting compound is soluble.
	When $\text{OH}^-$ pairs with $\text{Ca}^{2+}$ , $\text{Sr}^{2+}$ , or $\text{Ba}^{2+}$ , the resulting compound is slightly soluble.
$\text{CO}_3^{2-}$ and $\text{PO}_4^{3-}$	When these ions pair with $\text{Li}^+$ , $\text{Na}^+$ , $\text{K}^+$ , or $\text{NH}_4^+$ , the resulting compounds are soluble.

# Model Rules Of Solubility In Aqueous Solutions Pogil

**G. T. Hefter, R. P. T. Tomkins**



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