

Name \_\_\_\_\_

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## Chemistry – Unit 6 Worksheet 2

Balance the following equations by inserting the proper coefficients.

1. \_\_\_\_  $\text{SO}_2$  + \_\_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_\_  $\text{SO}_3$
2. \_\_\_\_  $\text{CH}_4$  + \_\_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_\_  $\text{CO}$  + \_\_\_\_  $\text{H}_2\text{O}$
3. \_\_\_\_  $\text{P}$  + \_\_\_\_  $\text{Cl}_2$   $\rightarrow$  \_\_\_\_  $\text{PCl}_3$
4. \_\_\_\_  $\text{CO}$  + \_\_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_\_  $\text{CO}_2$
5. \_\_\_\_  $\text{CH}_4$  + \_\_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_\_  $\text{CH}_3\text{OH}$
6. \_\_\_\_  $\text{Li}$  + \_\_\_\_  $\text{Br}_2$   $\rightarrow$  \_\_\_\_  $\text{LiBr}$
7. \_\_\_\_  $\text{Al}_2\text{O}_3$   $\rightarrow$  \_\_\_\_  $\text{Al}$  + \_\_\_\_  $\text{O}_2$
8. \_\_\_\_  $\text{Na}$  + \_\_\_\_  $\text{H}_2\text{O}$   $\rightarrow$  \_\_\_\_  $\text{NaOH}$  + \_\_\_\_  $\text{H}_2$
9. \_\_\_\_  $\text{CO}_2$  + \_\_\_\_  $\text{H}_2\text{O}$   $\rightarrow$  \_\_\_\_  $\text{C}_6\text{H}_{12}\text{O}_6$  + \_\_\_\_  $\text{O}_2$
10. \_\_\_\_  $\text{H}_2\text{SO}_4$  + \_\_\_\_  $\text{NaCl}$   $\rightarrow$  \_\_\_\_  $\text{HCl}$  + \_\_\_\_  $\text{Na}_2\text{SO}_4$

Two reactions used to get rid of sulfur dioxide, a pollutant from burning coal:

11. \_\_\_\_  $\text{H}_2$  + \_\_\_\_  $\text{SO}_2$   $\rightarrow$  \_\_\_\_  $\text{H}_2\text{S}$  + \_\_\_\_  $\text{H}_2\text{O}$
12. \_\_\_\_  $\text{CaCO}_3$  + \_\_\_\_  $\text{SO}_2$  + \_\_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_\_  $\text{CaSO}_4$  + \_\_\_\_  $\text{CO}_2$
13. \_\_\_\_  $\text{AgNO}_3$  + \_\_\_\_  $\text{CaCl}_2$   $\rightarrow$  \_\_\_\_  $\text{AgCl}$  + \_\_\_\_  $\text{Ca}(\text{NO}_3)_2$
14. \_\_\_\_  $\text{HCl}$  + \_\_\_\_  $\text{Ba}(\text{OH})_2$   $\rightarrow$  \_\_\_\_  $\text{BaCl}_2$  + \_\_\_\_  $\text{H}_2\text{O}$
15. \_\_\_\_  $\text{H}_3\text{PO}_4$  + \_\_\_\_  $\text{NaOH}$   $\rightarrow$  \_\_\_\_  $\text{Na}_3\text{PO}_4$  + \_\_\_\_  $\text{H}_2\text{O}$
16. \_\_\_\_  $\text{Pb}(\text{NO}_3)_2$  + \_\_\_\_  $\text{KI}$   $\rightarrow$  \_\_\_\_  $\text{PbI}_2$  + \_\_\_\_  $\text{KNO}_3$
17. \_\_\_\_  $\text{CuO}$  + \_\_\_\_  $\text{NH}_3$   $\rightarrow$  \_\_\_\_  $\text{N}_2$  + \_\_\_\_  $\text{Cu}$  + \_\_\_\_  $\text{H}_2\text{O}$
18. \_\_\_\_  $\text{C}_2\text{H}_5\text{OH}$  + \_\_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_\_  $\text{CO}_2$  + \_\_\_\_  $\text{H}_2\text{O}$
19. \_\_\_\_  $\text{C}_2\text{H}_6$  + \_\_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_\_  $\text{CH}_3\text{COOH}$  + \_\_\_\_  $\text{H}_2\text{O}$
20. \_\_\_\_  $\text{NO}_2$  + \_\_\_\_  $\text{H}_2\text{O}$   $\rightarrow$  \_\_\_\_  $\text{HNO}_3$  + \_\_\_\_  $\text{NO}$

Write the formulas of the reactants and products - including the symbols for the state - (s), (l), (g), (aq) - then balance the equations.

1. Gaseous ammonia ( $\text{NH}_3$ ) reacts with hydrogen chloride gas to form solid ammonium chloride.

# Modeling Chemistry U5 Ws1 V1 1 Key

**Yijin Wang**



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