

Chemistry – Unit 5 Worksheet 1

Relative Abundance of the Elements

Relative Mass From Counts

We can estimate relative abundances by counting all particles in the substance of pure composition and then:

1. Express abundance of gases as mass fractions of molecules or the gases themselves (molecules).
2. Express mass fractions as percentages (mass given as parts of 100).

Example: H_2 , O_2 , N_2 , Ar , CO_2 , H_2O

Example: CO_2 is 0.03% of the air. Expressing abundance here that the density of oxygen gas is expressed as percentage and percentage is 0.21 g/l. Expressing abundance as percentage, abundance is 0.21 g/l. Expressing abundance as percentage, abundance is 0.21 g/l. Expressing abundance as percentage, abundance is 0.21 g/l.

Oxygen
 2 g of O_2
 1 liter

Hydrogen
 1 g of H_2
 1 liter

Carbon
 1 g of C
 1 liter

Relative
 Density = $\frac{1.43 \text{ g O}_2}{1.29 \text{ g H}_2}$
 1.11 g/l
 1.11 g/l
 1.11 g/l

Expressing abundance as percentage, abundance is 0.21 g/l. Expressing abundance as percentage, abundance is 0.21 g/l. Expressing abundance as percentage, abundance is 0.21 g/l. Expressing abundance as percentage, abundance is 0.21 g/l.

Relative Mass From Composition

Many substances are composed of elements. Expressing abundance as percentage, abundance is 0.21 g/l. Expressing abundance as percentage, abundance is 0.21 g/l. Expressing abundance as percentage, abundance is 0.21 g/l.

Example: CO_2 is 0.03% of the air. Expressing abundance as percentage, abundance is 0.21 g/l. Expressing abundance as percentage, abundance is 0.21 g/l. Expressing abundance as percentage, abundance is 0.21 g/l.



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