Name:	Date:Pd:
	Chemical Reactions CTICE TEST/ REVIEW
A chemical reaction is a process in wh	iich
a. products change into reactants	c. substances can change state
b. the law of conservation of mass applies	d. all of these
2. During a chemical reaction,	
a. new elements are produced	c. atoms are destroyed
b. atoms are rearranged	d. elements are destroyed
3. An equation is balanced by	
a. changing subscripts	c. erasing elements as necessary
b. adding coefficients	d. adding elements as necessary
4. What are the reactants in the following	g chemical equation:
$Zn + CuSO_4 \rightarrow ZnSO_4 + Cu$	
a. zinc and copper	c. zinc and copper (II) sulfate
b. zinc sulfate and copper	d. only zinc
5. What are the products in the above equ	uation?
a. zinc and copper	c. zinc and copper (II) sulfate
b. zinc sulfate and copper	d. only zinc
6. The breaking down of a substance into	two or simpler substances is
Displacement Synthesis	c. Decomposition d. Double displacement
7. Each substance to the right of the arro-	w in a chemical equation is a(n)
a. Reactant	e. Product
b. Catalyst	d. Inhibitor
According to the law of conservation of many atoms of hydrogen must be part of the present the present that the present the present that the present that the present the present that the present the present the present the present the present that the present the	of mass, if two atoms of hydrogen are used as a reactant, how oduct?
a. 4	c. 1
b. 2	d. 0
9. What type of reaction is shown in the	following chemical equation, $NH_4C1 \rightarrow NH_3 + HC1$?
a. Displacement	c. Decomposition
b. Synthesis	d. Double displacement

Modeling Chemistry Unit 7 Chemical Reactions Answer Key

Michael Brown

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