

Chemistry Unit 8 Worksheet 4

Samples of Every Kind of Problem

On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedure outlined in class. Be sure to circle your final answer.

1. Calculate the number of moles of potassium chlorate, KClO_3 (s), that must decompose to produce potassium chloride, KCl (s), and 1.8 moles of oxygen gas.

| | | | | |
|-----------|---|------|------|-------------------------------------|
| Equation: | $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$ | | | 1.2 moles KClO_3 needed |
| Before: | 1.2 | 0 | 0 | |
| Change: | -1.2 | +1.2 | +1.8 | |
| After: | 0 | 1.2 | 1.8 | |

2. In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 2.43 g of magnesium?

| | | | | | |
|-----------|---|----------|---------|-------|------------------------------------|
| Equation: | $\text{Mg} + 2 \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$ | | | | 0.200 moles HCl needed |
| Before: | 0.100 | ? | 0 | 0 | |
| Change: | -0.100 | -0.200 + | 0.100 + | 0.100 | |
| After: | 0 | ? | 0.100 | 0.100 | |

3. Ethane, C_2H_6 reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.20 moles of ethane?

| | | | | |
|-----------|--|--------|--------|--------|
| Equation: | $2 \text{C}_2\text{H}_6 + 7 \text{O}_2 \rightarrow 4 \text{CO}_2 + 6 \text{H}_2\text{O}$ | | | |
| Before: | 2.20 | xs | 0 | 0 |
| Change: | -2.20 | - 7.70 | + 4.40 | + 6.60 |
| After: | 0 | xs | 4.40 | 6.60 |

$$\text{O}_2 \text{ needed: } 7.70 \text{ mole} \times \frac{32.0 \text{ g}}{1 \text{ mole}} = 246 \text{ g O}_2 \text{ required}$$

4. Determine the mass of sodium nitrate produced when 0.73 g of nickel (II) nitrate reacts with sodium hydroxide according to the following equation:



$$0.73 \text{ g} \times \frac{1 \text{ mole}}{183 \text{ g}} = .0040 \text{ mole}$$

| | | | | |
|-----------|---|----------|---------|---------|
| Equation: | $\text{Ni}(\text{NO}_3)_2 + 2 \text{NaOH} \rightarrow \text{Ni}(\text{OH})_2 + 2 \text{NaNO}_3$ | | | |
| Before: | 0.0040 | XS | 0 | 0 |
| Change: | - 0.0040 | - 0.0080 | +0.0040 | +0.0080 |
| After: | 0 | XS | 0.0040 | 0.0080 |

$$\text{NaNO}_3 \text{ produced: } 0.0080 \text{ mole} \times \frac{85.0 \text{ g}}{1 \text{ mole}} = 0.68 \text{ g}$$

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