Practice MCAT Test I

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Passage I

Half reaction

Below is a table showing some Standard State Reduction Potentials.

E. OVA

The state of the s	100000000000000000000000000000000000000
$F_* + 2e^* \rightarrow 2F^*$	2.87
$Cl_x + 2e^x \rightarrow 2Cl^x$	1.36
$Br_* + 2e^* \rightarrow 2Br^*$	1.09
$Ag^* + e^* \rightarrow Ag$	0.80
$Fe^{y_1} + e^y \rightarrow Fe^{y_2}$	0.77
$Cu^{2+} + 2e^{-} \rightarrow Cu$	0.34
$2H^* + 2e^* \rightarrow H_*$	0.00
Pb ²⁺ + 2e ⁻ → P6	-0.13
$Ni^{2*} + 2e^* \rightarrow Ni$	-0.25
$Zn^{2+} + 2e^- \rightarrow Zn$	-0.76
$Na^* + e^* \rightarrow Na$	-2.71

- What should happen when a piece of copper is placed in 1M HC1?
 - The copper is dissolved by the acid.
 - B. The copper is dissolved by the acid with the release of hydrogen gas.
 - C. The copper bursts into greenish flames.
 - D. Nothing happens.
- What should happen when a piece of lead is placed in 1M HC1?
 - A. The lead is dissolved by the acid.
 - B. The lead begins to dissolve with the release of hydrogen gas.
- C. The lead bursts into flames.
- D. Nothing happens.
- How many voltaic cells with a voltage greater than 2V can be made using the standard half-cell reactions histod?
 - A . 11
 - B. 9
 - C. 17
 - D. 15
- Suppose that the half-cell reaction Ni2+ 2e → Ni is assigned to have a potential of 0.00 V The standard reduction potential for the reduction of Br, would be? A. 1.34 V B. 0.84 V
 - C. 1.09 V D. 1.25 V

- A voltaic cell is set up with F/F as one half cell and Br/ Br as the other. What is the voltage of the cell at standard state?
 - A. 3.96 V B. 1.78 V.
 - C. 1.09 V
 - D. 2.87 V
- NaCl is melted in a beaker and electrolysis carried out. What is the minimum potential difference required to separate sodium and chlorine?
 - A. -4.07 V
 - B. -1.36 V
 - C. 1.36 V D. 4.07 V
- If standard state oxidation potentials are used instead of standard state reduction potentials and the half reactions are listed in descending order, which one of the following would be true?
 - A. the lead reaction would be above the silver reaction
 - B. the fluorine reaction would be above the chlorine reaction
 - the iron reaction would be above the nickel reaction.
 - the hydrogen reaction would be below the copper reaction

Passage III

A student is provided with a variety of electrical components and asked to perform certain tasks.

- She is provided with three 2 resistors. How many different total resistances can be produced if all the resistors must be used in a circuit?
 - A. 3
 - B. 4
 - C. 5
 - D. 6
- What are the minimum and maximum total resistances produced above?
 - A. 0.67 and 3
 - B. 1 and 6
 - C. 1.33 and 6
 - D. 0.67 and 6
- What is the total resistance of two 2 resistors in parallel. connected to a third 2 resistor in series?
 - A. 6 B. 1
 - C. 3
 - D. 1.33

Mcat Past Papers With Answers

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