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CHAPTER 12**STUDY GUIDE FOR CONTENT MASTERY****Section 12.3 Limiting Reactants**

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 15–17 in the space below the table.



The molar masses of O_2 , NO , and NO_2 are 32.00 g/mol, 30.01 g/mol, and 46.01 g/mol, respectively.

Amount of O_2	Amount of NO	Amount of NO_2	Limiting Reactant	Amount and Mass of Excess Reactant
1 molecule	2 molecules	2 molecules	none	none
2 molecules	4 molecules	4 molecules	NO	2 molecules O_2
2 molecules	4 molecules	4 molecules	1. O_2	2. 4 molecules NO
1.00 mol	2.00 mol	2.00 mol	3. none	4. none
4.00 mol	6.00 mol	4.00 mol	5. NO	6. 2.00 mol O_2
6.00 mol	12.00 mol	6.00 mol	7. NO	8. 1.50 mol O_2
1.00 mol	6.00 mol	10.00 mol	9. O_2	10. 5.00 mol NO
6.000 mol	12.000 mol	16.000 mol	11. NO	12. 6.000 mol O_2
32.00 g	60.02 g	16.02 g	13. none	14. none
64.00 g	60.02 g	16.02 g	15. O_2	16. 30.12 g NO
64.00 g	120.04 g	32.04 g	17. O_2	18. 1.24 g NO

balanced equation mole ratio = 2 mol NO /1 mol O_2

$16.00 \text{ g } \text{O}_2 \times 1 \text{ mol } \text{O}_2 / 32.00 \text{ g } \text{O}_2 = 0.5125 \text{ mol } \text{O}_2$

$20.00 \text{ g } \text{NO} \times 1 \text{ mol } \text{NO} / 30.01 \text{ g } \text{NO} = 0.6664 \text{ mol } \text{NO}$

actual mole ratio = $0.6664 \text{ mol } \text{NO} / 0.5125 \text{ mol } \text{O}_2 = 1.302 \text{ mol } \text{NO} / 1.000 \text{ mol } \text{O}_2$

Because the actual mole ratio of $\text{NO}:\text{O}_2$ is larger than the balanced equation mole ratio of $\text{NO}:\text{O}_2$, there is an excess of NO ; O_2 is the limiting reactant.

Mass of NO used = $0.5125 \text{ mol } \text{O}_2 \times 2 \text{ mol } \text{NO} / 1 \text{ mol } \text{O}_2 = 0.4250 \text{ mol } \text{NO}$

$0.4250 \text{ mol } \text{NO} \times 30.01 \text{ g } \text{NO} / 1 \text{ mol } \text{NO} = 12.76 \text{ g } \text{NO}$

Mass of NO_2 produced = $0.4250 \text{ mol } \text{NO}_2 \times 46.01 \text{ g } \text{NO}_2 / 1 \text{ mol } \text{NO}_2 = 19.56 \text{ g } \text{NO}_2$

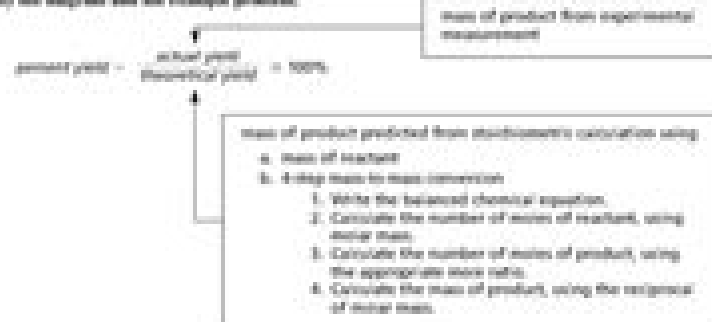
Excess $\text{NO} = 20.00 \text{ g } \text{NO} - 12.76 \text{ g } \text{NO} = 7.24 \text{ g } \text{NO}$

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CHAPTER 13**STUDY GUIDE FOR CONTENT MASTERY****Section 13.4 Percent Yield**

In your textbook, read about the yields of products.

Study the diagram and the example problem.



Example Problem: The following chemical equation represents the production of palladium, a valuable metal in the manufacturing of some semiconductor devices.



In one experiment, the reaction yielded 7.42 g of the metal from a 7.00-g sample of palladium. Determine the percent yield of this reaction. The molar masses of Ga and Ga_2O_3 are 69.72 g/mol and 157.44 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

1. actual yield unknown 7.42 g Ga_2O_3
2. mass of reactant: 7.00 g Ga correct
3. number of moles of reactant: $7.00 \text{ g } \text{Ga} \times \frac{69.72 \text{ g } \text{Ga}}{1 \text{ mol } \text{Ga}} = 1 \text{ mol } \text{Ga}$ 7.00 g $\text{Ga} \times 1 \text{ mol } \text{Ga} / 69.72 \text{ g } \text{Ga}$
4. number of moles of product: $0.100 \text{ mol } \text{Ga} \times \frac{2 \text{ mol } (\text{Ga}_2\text{O}_3)}{1 \text{ mol } \text{Ga}} = 0.200 \text{ mol } \text{Ga}_2\text{O}_3$ 0.100 mol $\text{Ga} \times \frac{2 \text{ mol } \text{Ga}_2\text{O}_3}{1 \text{ mol } \text{Ga}}$
5. theoretical yield: $0.200 \text{ mol } (\text{Ga}_2\text{O}_3) \times \frac{157.44 \text{ g } (\text{Ga}_2\text{O}_3)}{1 \text{ mol } (\text{Ga}_2\text{O}_3)} = 31.49 \text{ g } \text{Ga}_2\text{O}_3$ correct
6. percent yield: $\frac{7.42 \text{ g } (\text{Ga}_2\text{O}_3)}{31.49 \text{ g } (\text{Ga}_2\text{O}_3)} \times 100 = 23.6\%$ $\frac{7.42 \text{ g } \text{Ga}_2\text{O}_3}{31.49 \text{ g } \text{Ga}_2\text{O}_3} \times 100$

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CH Cherryholmes



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akan otomatis terhitung contoh penggunaan rumus excel total harga

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web rumus penjumlahan excel ke samping sebenarnya adalah penggunaan autofill pada rumus jumlah otomatis sum untuk menghitung total kolom lainnya pada tabel contoh menghitung jumlah total penjualan di provinsi jawa timur dan total keseluruhan

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web pada kolom g baris 3 ketik simbol sebagai pengenalan rumus selanjutnya kita akan menggunakan perkalian agar dapat diketahui total harganya yang akan dikalikan adalah kolom satuan dan kolom harga satuan sehingga rumusnya seperti berikut a angka langsung 10 3000

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web nov 4 2021 count adalah rumus yang digunakan untuk menghitung jumlah cell excel yang kamu pilih caranya gampang tinggal kalikan harga item x jumlah barang atau cell b2 x c2 cara menghitungnya yaitu harga asli x harga diskon atau a2 b2

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web anda bisa merujuk ke nilai yang ditampilkan pada bilah status saat anda ingin sekilas melihat data namun tidak memiliki waktu untuk memasukkan rumus video menghitung sel menggunakan bilah status excel tonton video berikut untuk mempelajari cara menampilkan jumlah pada bilah status