

Name _____ Date _____ Class _____

CHAPTER 12**STUDY GUIDE FOR CONTENT MASTERY****Section 12.3 Limiting Reactants**

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 25–27 in the space below the table.



The molar masses of O_2 , N_2 , and NO_2 are 32.00 g/mol, 28.02 g/mol, and 46.01 g/mol, respectively.

Amount of O_2	Amount of N_2	Amount of NO_2	Limiting Reactant	Amount and Mass of Excess Reactant
1 molecule	2 molecules	2 molecules	none	none
4 molecules	4 molecules	4 molecules	N_2	2 molecules O_2
2 molecules	4 molecules	4 molecules	O_2	2 molecules NO_2
1.00 mol	2.00 mol	2.00 mol	none	none
4.00 mol	4.00 mol	4.00 mol	N_2	2.00 mol O_2
2.00 mol	4.00 mol	4.00 mol	O_2	2.00 mol NO_2
1.00 mol	4.00 mol	4.00 mol	O_2	2.00 mol NO_2
0.500 mol	2.00 mol	2.00 mol	N_2	0.500 mol O_2
22.00 g	46.01 g	46.01 g	none	none
44.00 g	46.01 g	46.01 g	O_2	10.12 g NO_2
44.00 g	28.02 g	28.02 g	O_2	1.24 g NO_2

balanced equation mole ratio = 2 mol NO_2 /1 mol O_2

16.00 g O_2 $\times$ 1 mol O_2 /32.00 g O_2 = 0.500 mol O_2

28.02 g N_2 $\times$ 1 mol N_2 /28.02 g N_2 = 0.999 mol N_2

actual mole ratio = 0.999 mol N_2 /0.500 mol O_2 = 2.00 mol N_2 /1.00 mol O_2

Because the actual mole ratio of N_2 / O_2 is larger than the balanced equation mole ratio of N_2 / O_2 , there is an excess of N_2 ; O_2 is the limiting reactant.

Mass of NO_2 used = 0.500 mol O_2 $\times$ 2 mol NO_2 /1 mol O_2 = 0.500 mol NO_2

0.500 mol NO_2 $\times$ 46.01 g NO_2 /1 mol NO_2 = 23.01 g NO_2

Mass of NO_2 produced = 0.500 mol NO_2 $\times$ 46.01 g NO_2 /1 mol NO_2 = 23.01 g NO_2

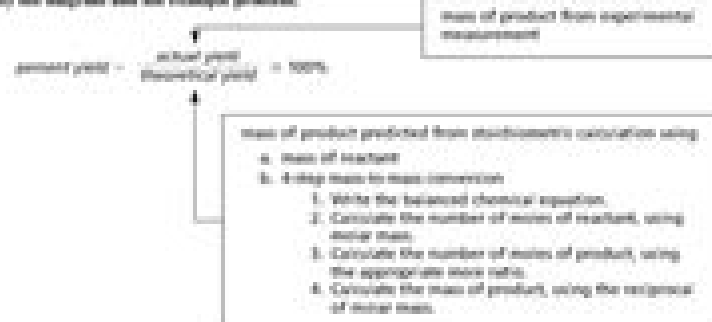
Excess N_2 = 28.02 g N_2 - 23.01 g N_2 = 5.01 g N_2

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CHAPTER 13**STUDY GUIDE FOR CONTENT MASTERY****Section 13.4 Percent Yield**

In your textbook, read about the yield of products.

Study the diagram and the example problem.



Example Problem: The following chemical equation represents the production of palladium, a valuable metal in the manufacturing of some semiconductor devices.



In one experiment, the reaction yielded 7.42 g of the metal from a 7.00-g sample of palladium. Determine the percent yield of this reaction. The molar masses of Ga and Ga_2O_3 are 69.72 g/mol and 157.44 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- actual yield: unknown 7.42 g Ga_2O_3
- mass of reactant: 7.00 g Ga correct
- number of moles of reactant: 7.00 g Ga $\times$ $\frac{69.72 \text{ g Ga}}{1 \text{ mol Ga}}$ = 1 mol Ga/69.72 g Ga
- number of moles of product: 0.100 mol Ga $\times$ $\frac{2 \text{ mol Ga}_2\text{O}_3}{1 \text{ mol Ga}}$ = 0.200 mol Ga_2O_3
- theoretical yield: 0.200 mol Ga_2O_3 $\times$ $\frac{157.44 \text{ g Ga}_2\text{O}_3}{1 \text{ mol Ga}_2\text{O}_3}$ = 31.49 g Ga_2O_3 correct
- percent yield: $\frac{7.42 \text{ g Ga}_2\text{O}_3}{31.49 \text{ g Ga}_2\text{O}_3} \times 100$ = 23.6% correct

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Table of Contents Mcgraw Hill Connect Chemistry Study Guide Answers

1. Understanding the eBook Mcgraw Hill Connect Chemistry Study Guide Answers
 - The Rise of Digital Reading Mcgraw Hill Connect Chemistry Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Mcgraw Hill Connect Chemistry Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mcgraw Hill Connect Chemistry Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mcgraw Hill Connect Chemistry Study Guide Answers
 - Personalized Recommendations
 - Mcgraw Hill Connect Chemistry Study Guide Answers User Reviews and Ratings
 - Mcgraw Hill Connect Chemistry Study Guide Answers and Bestseller Lists

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8. Staying Engaged with Mcgraw Hill Connect Chemistry Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mcgraw Hill Connect Chemistry Study Guide Answers
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 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mcgraw Hill Connect Chemistry Study Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - Carving Out Dedicated Reading Time
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 - Fact-Checking eBook Content of Mcgraw Hill Connect Chemistry Study Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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