

Mixed Nomenclature

Identify each compound as Ionic, Covalent, or Acid, then complete the missing information.

	Type of Compound (Ionic, Covalent, or Acid)	Name	Chemical Formula
1.	Ionic	⁺² Magnesium sulfide ⁻²	MgS
2.	Covalent	Tetrasulfur octaoxide	S ₄ O ₈
3.	Ionic	Magnesium chloride	MgCl ₂
4.	Ionic	⁺¹ Cesium carbonate ⁻²	Cs ₂ CO ₃
5.	Acid	Hydrofluoric acid	HF
6.	Covalent	Octanitrogen hexaoxide	N ₈ O ₆
7.	Ionic	⁺² Manganese (II) chloride ⁻¹	MnCl ₂
8.	Acid	Phosphoric acid	H ₃ PO ₄
9.	Ionic	⁺¹ Ammonium sulfate ⁻²	(NH ₄) ₂ SO ₄
10.	Ionic	Iron (III) sulfide	Fe ₂ S ₃
11.	Acid	Chloric acid	HClO ₃
12.	Ionic	⁺³ Aluminum chlorate ⁻¹	Al(ClO ₃) ₃
13.	Covalent	Dinitrogen trioxide	N ₂ O ₃
14.	Ionic	⁺¹ Ammonium nitrite ⁻¹	NH ₄ NO ₂
15.	Ionic	Sodium nitride	Na ₃ N
16.	Covalent	Trisulfur nonafluoride	S ₃ F ₉
17.	Ionic	Copper (II) acetate	Cu(C ₂ H ₃ O ₂) ₂
18.	Ionic	Calcium phosphate	Ca ₃ (PO ₄) ₂
19.	Acid	Nitrous acid	HNO ₂
20.	Ionic	⁺² Beryllium cyanide ⁻¹	Be(CN) ₂
21.	Covalent	Phosphorus pentabromide	PBr ₅
22.	Ionic	⁺⁴ Tin (IV) carbonate ⁻²	Sn(CO ₃) ₂
23.	Acid	Hydronitric acid	H ₃ N
24.	Ionic	Nickel (II) sulfide	NiS
25.	Covalent	Tetrachlorine heptabromide	Cl ₄ Br ₇

Mixed Nomenclature And Properties Worksheet

JE Gale



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