		Gas Law Formula		
Gas Law	Formula	Description		
Boyle's Law	$P_1V_1 = P_2V_2$	At constant T, as pressure increases, volume decreases.		
Charles' Law	$\frac{V_1}{T_1} = \frac{V_2}{T_2}$	At constant P, as volume increases, temperature increases.		
Gay-Lussac's Law	$\frac{P_1}{T_1} = \frac{P_2}{T_2}$	At constant V, as pressure increases, temperature increases.		
Combined Law	$\frac{P_1V_1}{T_1} = \frac{P_2V}{T_2}$	Obtained by combining E Lussac's Law.	Obtained by combining Boyle's Law, Charles' Law and Gay- Lussac's Law.	
Ideal Gas Law	PV = nRT	• •		
		P = pressure in kPa n = number of moles	R = ideal gas constant	

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