

# Oxidation and Reduction

What happens when electrons are transferred in a chemical reaction?

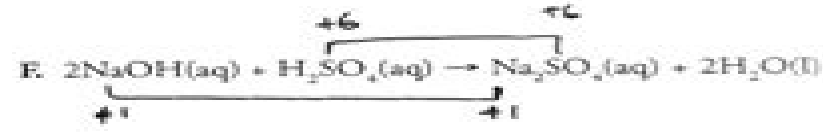
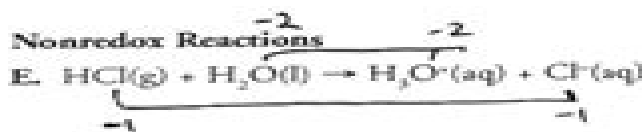
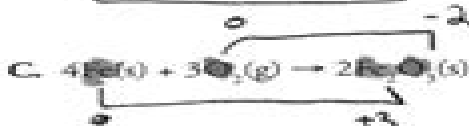
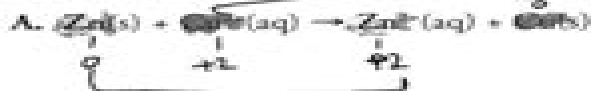
KEY

## Why?

Silver tarnishes when it comes in contact with sulfur compounds in the air. Copper gets coated in beautiful green patina as it ages. Metals rust or corrode in the presence of air and water. Minerals (ionic compounds) found in ore can be decomposed with the use of electricity to produce pure metals and nonmetals. All of these reactions are examples of oxidation and reduction, otherwise known as **redox** reactions. In this activity you will explore what is happening at the atomic level in redox reactions.

## Model 1 – Redox Reactions

### Redox Reactions



1. What two types of reactions are shown in Model 1?

*Redox & Nonredox*

2. Examine the redox and nonredox reactions in Model 1. Is/are there any feature(s) in the redox reactions that would allow you to identify them as redox reactions? If yes, use specific examples from Model 1 to support your answer.

*Not necessarily*

# Model 1 Redox Reactions Answers Pogil

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Ekborg, Anders Berg**



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## **Table of Contents Model 1 Redox Reactions Answers Pogil**

1. Understanding the eBook Model 1 Redox Reactions Answers Pogil
  - The Rise of Digital Reading Model 1 Redox Reactions Answers Pogil
  - Advantages of eBooks Over Traditional Books
2. Identifying Model 1 Redox Reactions Answers Pogil
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Model 1 Redox Reactions Answers Pogil
  - User-Friendly Interface
4. Exploring eBook Recommendations from Model 1 Redox Reactions Answers Pogil
  - Personalized Recommendations
  - Model 1 Redox Reactions Answers Pogil User Reviews and Ratings
  - Model 1 Redox Reactions Answers Pogil and Bestseller Lists
5. Accessing Model 1 Redox Reactions Answers Pogil Free and Paid eBooks
  - Model 1 Redox Reactions Answers Pogil Public Domain eBooks
  - Model 1 Redox Reactions Answers Pogil eBook Subscription Services
  - Model 1 Redox Reactions Answers Pogil Budget-Friendly Options
6. Navigating Model 1 Redox Reactions Answers Pogil eBook Formats

- ePub, PDF, MOBI, and More
  - Model 1 Redox Reactions Answers Pogil Compatibility with Devices
  - Model 1 Redox Reactions Answers Pogil Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Model 1 Redox Reactions Answers Pogil
  - Highlighting and Note-Taking Model 1 Redox Reactions Answers Pogil
  - Interactive Elements Model 1 Redox Reactions Answers Pogil
8. Staying Engaged with Model 1 Redox Reactions Answers Pogil
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Model 1 Redox Reactions Answers Pogil
9. Balancing eBooks and Physical Books Model 1 Redox Reactions Answers Pogil
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Model 1 Redox Reactions Answers Pogil
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Model 1 Redox Reactions Answers Pogil
- Setting Reading Goals Model 1 Redox Reactions Answers Pogil
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Model 1 Redox Reactions Answers Pogil
- Fact-Checking eBook Content of Model 1 Redox Reactions Answers Pogil
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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