

## Chemistry – Unit 5 Worksheet 1

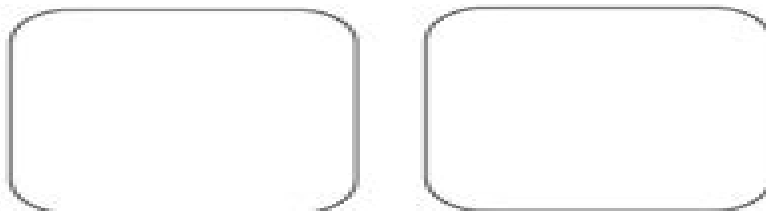
### Molar Masses of the Elements

#### Relative Mass From Gases

We have established that the combining ratio of gases can be explained if two assumptions are made:

1. Equal volumes of gases contain the same number of molecules at the same pressure and temperature.
2. Some pure elemental gases are clustered into pairs to form diatomic molecules.

Use these assumptions and particle diagrams to explain the fact that the density of oxygen gas at standard temperature and pressure is 1.43 g/liter, whereas the density of hydrogen gas under these conditions is 0.089 g/liter. How many times more massive is one molecule of oxygen than one molecule of hydrogen?



You shouldn't conclude that chemists were able to determine the molar masses of all the elements using this technique. Measurements of the density of the gaseous phase of many of the elements are difficult, if not impossible. However, we are going to see that chemists could use another tool – the *percent composition of compounds* to determine molar masses.

#### Relative Mass From Compounds

Many substances combine with oxygen to form a type of compound called an oxide. In the previous unit we saw that such combinations often occur in multiple proportions. John Dalton made the assumption that the lowest ratio was a 1:1 combination of elements. For now, we will make a similar assumption. *We may have to re-examine this assumption later.*

In Unit 4 you examined % composition data of several compounds. For example, 100 g of the simplest oxide of carbon contains 42.9g of C and 57.1 g of O. To be able to compare masses of elements to the mass of oxygen, it is important to use the same mass of oxygen in all of the ratios. Using a proportion, one can show that 75.0 g of C combines with 100 g of oxygen.

$$\frac{42.9\text{gC}}{57.1\text{gO}} = \frac{75.0\text{gC}}{100\text{gO}}$$

# Modeling Chemistry Unit 1

**TD Snyder**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

## **Modeling Chemistry Unit 1:**

## Unveiling the Magic of Words: A Overview of "**Modeling Chemistry Unit 1**"

In a world defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is truly awe-inspiring. Enter the realm of "**Modeling Chemistry Unit 1**," a mesmerizing literary masterpiece penned with a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve to the book is central themes, examine its distinctive writing style, and assess its profound impact on the souls of its readers.

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