

Chemistry – Unit 6 Notes 1 - Thomson Model of the Atom

J. J. Thomson performed experiments with cathode rays in an attempt to understand electricity – which was still a mystery in the late 1800s. Use the website *A Look Inside the Atom*¹ to find the conclusions that Thomson and other physicists drew regarding the mysterious cathode rays.

Thomson's 1897 Experiments - state the conclusions Thomson drew from each of his famous cathode ray experiments:

1. **First Experiment:** Thomson directed the beam at an electrometer and tried to separate the evidence of charge from the path of the beam. *What connection did Thomson find between charge and the cathode rays? Was the charge positive or negative?*
2. **Second Experiment:** Thomson tried passing the cathode ray through an electric field. *How did cathode ray beam behave when it passed through an electric field? What did he conclude after his second experiment?*
3. **Third Experiment:** Thomson did some careful measurements on how much the path of the cathode ray was bent in a magnetic field and how much energy they carried. From this work Thomson could describe the mass/charge ratio of the cathode ray particles. *What amazing result did Thomson find?*

Thomson's Atomic Model: Thomson presented three hypotheses from his experiments. Only two were accepted by physicists – in fact the third was shown to be wrong! From the first two came a model of the atom known as the *Plum Pudding* model. Complete the atom drawing below to illustrate Thomson's plum pudding model. Explain how this fits with his observations.



¹ <https://www.aip.org/history/exhibits/electron/jjrays.htm>
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Modeling Chemistry U6 Ws 4 V0

Lauren Gardner



Modeling Chemistry U6 Ws 4 V0:

Whispering the Strategies of Language: An Emotional Quest through **Modeling Chemistry U6 Ws 4 V0**

In a digitally-driven world wherever displays reign supreme and quick connection drowns out the subtleties of language, the profound strategies and emotional subtleties hidden within phrases frequently get unheard. However, located within the pages of **Modeling Chemistry U6 Ws 4 V0** a interesting fictional prize pulsating with natural thoughts, lies an exceptional quest waiting to be undertaken. Penned by an experienced wordsmith, that enchanting opus invites readers on an introspective trip, delicately unraveling the veiled truths and profound affect resonating within the cloth of each word. Within the mental depths of the poignant evaluation, we shall embark upon a sincere exploration of the book is core styles, dissect their captivating publishing model, and yield to the powerful resonance it evokes strong within the recesses of readers hearts.

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