

## Chemistry Unit 8 Worksheet 4

### Samples of Every Kind of Problem

*On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedure outlined in class. Be sure to circle your final answer.*

1. Calculate the number of moles of potassium chlorate,  $\text{KClO}_3$  (s), that must decompose to produce potassium chloride,  $\text{KCl}$  (s), and 1.8 moles of oxygen gas.

Equation:	$2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$			1.2 moles $\text{KClO}_3$ needed
Before:	1.2	0	0	
Change:	-1.2	+1.2	+1.8	
After:	0	1.2	1.8	

2. In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 2.43 g of magnesium?

Equation:	$\text{Mg} + 2 \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$				0.200 moles $\text{HCl}$ needed
Before:	0.100	?	0	0	
Change:	-0.100	-0.200 +	0.100 +	0.100	
After:	0	?	0.100	0.100	

3. Ethane,  $\text{C}_2\text{H}_6$  reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.20 moles of ethane?

Equation:	$2 \text{C}_2\text{H}_6 + 7 \text{O}_2 \rightarrow 4 \text{CO}_2 + 6 \text{H}_2\text{O}$			
Before:	2.20	xs	0	0
Change:	-2.20	-7.70	+4.40	+6.60
After:	0	xs	4.40	6.60

$$\text{O}_2 \text{ needed: } 7.70 \text{ mole} \times \frac{32.0 \text{ g}}{1 \text{ mole}} = 246 \text{ g O}_2 \text{ required}$$

4. Determine the mass of sodium nitrate produced when 0.73 g of nickel (II) nitrate reacts with sodium hydroxide according to the following equation:



$$0.73 \text{ g} \times \frac{1 \text{ mole}}{183 \text{ g}} = .0040 \text{ mole}$$

Equation:	$\text{Ni}(\text{NO}_3)_2 + 2 \text{NaOH} \rightarrow \text{Ni}(\text{OH})_2 + 2 \text{NaNO}_3$			
Before:	0.0040	XS	0	0
Change:	-0.0040	-0.0080	+0.0040	+0.0080
After:	0	XS	0.0040	0.0080

$$\text{NaNO}_3 \text{ produced: } 0.0080 \text{ mole} \times \frac{85.0 \text{ g}}{1 \text{ mole}} = 0.68 \text{ g}$$

# Modeling Chemistry Unit 4

**Frank M. Dunnivant, Elliot Anders**



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## **Table of Contents Modeling Chemistry Unit 4**

1. Understanding the eBook Modeling Chemistry Unit 4
  - The Rise of Digital Reading Modeling Chemistry Unit 4
  - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry Unit 4
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modeling Chemistry Unit 4
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry Unit 4
  - Personalized Recommendations
  - Modeling Chemistry Unit 4 User Reviews and Ratings
  - Modeling Chemistry Unit 4 and Bestseller Lists
5. Accessing Modeling Chemistry Unit 4 Free and Paid eBooks
  - Modeling Chemistry Unit 4 Public Domain eBooks
  - Modeling Chemistry Unit 4 eBook Subscription Services
  - Modeling Chemistry Unit 4 Budget-Friendly Options

6. Navigating Modeling Chemistry Unit 4 eBook Formats
  - ePub, PDF, MOBI, and More
  - Modeling Chemistry Unit 4 Compatibility with Devices
  - Modeling Chemistry Unit 4 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Modeling Chemistry Unit 4
  - Highlighting and Note-Taking Modeling Chemistry Unit 4
  - Interactive Elements Modeling Chemistry Unit 4
8. Staying Engaged with Modeling Chemistry Unit 4
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Modeling Chemistry Unit 4
9. Balancing eBooks and Physical Books Modeling Chemistry Unit 4
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Modeling Chemistry Unit 4
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Modeling Chemistry Unit 4
  - Setting Reading Goals Modeling Chemistry Unit 4
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modeling Chemistry Unit 4
  - Fact-Checking eBook Content of Modeling Chemistry Unit 4
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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