

7 • Molecular Structure

STATION 3 • POLARITY

Determine the shape of the molecule and decide whether each compound is polar or nonpolar.

PH_3 $3 \times 1 = 3$ $\text{H} : \underset{\cdot\cdot}{\underset{\cdot\cdot}{\text{P}}} : \text{H}$ Polar	CF_4 $4 \times 4 = 16$ $\begin{array}{c} :\ddot{\text{F}}: \\ :\ddot{\text{F}}: \text{C} : \ddot{\text{F}}: \\ :\ddot{\text{F}}: \end{array}$ non-polar	BH_3 $3 \times 3 = 6$ $\text{H} : \underset{\cdot\cdot}{\underset{\cdot\cdot}{\text{B}}} : \text{H}$ non-polar
SO_2 6 $\begin{array}{c} :\ddot{\text{O}}: \text{S} :\ddot{\text{O}}: \\ \text{or } :\ddot{\text{O}}: \text{S} :\ddot{\text{O}}: \\ \text{or } :\ddot{\text{O}}: \text{S} :\ddot{\text{O}}: \end{array}$ Polar	CO_2 $12 \times 4 = 16$ $\begin{array}{c} :\ddot{\text{O}}: \text{C} :\ddot{\text{O}}: \\ \text{or } :\ddot{\text{O}}: \text{C} :\ddot{\text{O}}: \end{array}$ non-polar	SCl_2 $6 + 6 = 12$ $\begin{array}{c} :\ddot{\text{Cl}}: \text{S} :\ddot{\text{Cl}}: \\ \text{or } :\ddot{\text{Cl}}: \text{S} :\ddot{\text{Cl}}: \end{array}$ polar

7 • Molecular Structure

STATION 4 • IMF's AND ATTRACTIONS

Match the statement with the Inter-Particle Forces

A. Dipole-Dipole Interactions

B. Covalent bonds

C. Hydrogen-bonding

D. Ionic Bonds

E. Metallic Bonds

F. London Dispersion Forces

- F 1. Is used to explain why BP of $\text{Xe} > \text{Kr} > \text{Ar} > \text{Ne} > \text{He}$
- B 2. Is present in ~~Complete~~ all ~~Compounds~~ Compounds
- E 3. Is used to explain why Cu is a good conductor
- D 4. Is used to explain why NaCl(s) is a good conductor
- A 5. Is used to explain why ICI has a higher BP than Br_2
- C 6. Is used to explain why boiling point of CH_3OH is higher than CH_4
- E 7. A lattice of positive ions in a sea of delocalized electrons
- A 8. Positive ends of molecules attract negative ends of other molecules
- E 9. Lattice of alternating positively and negatively charged particles
- C 10. Hydrogen atoms tethered between small, highly electronegative atoms N, O, and F.
- F 11. Polarized electron clouds induce dipoles in their neighboring atoms

Modeling Chemistry Unit 7 Review Answers

Richard Bailey



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Table of Contents Modeling Chemistry Unit 7 Review Answers

1. Understanding the eBook Modeling Chemistry Unit 7 Review Answers
 - The Rise of Digital Reading Modeling Chemistry Unit 7 Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry Unit 7 Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Chemistry Unit 7 Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry Unit 7 Review Answers
 - Personalized Recommendations
 - Modeling Chemistry Unit 7 Review Answers User Reviews and Ratings
 - Modeling Chemistry Unit 7 Review Answers and Bestseller Lists
5. Accessing Modeling Chemistry Unit 7 Review Answers Free and Paid eBooks
 - Modeling Chemistry Unit 7 Review Answers Public Domain eBooks
 - Modeling Chemistry Unit 7 Review Answers eBook Subscription Services
 - Modeling Chemistry Unit 7 Review Answers Budget-Friendly Options

6. Navigating Modeling Chemistry Unit 7 Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Modeling Chemistry Unit 7 Review Answers Compatibility with Devices
 - Modeling Chemistry Unit 7 Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modeling Chemistry Unit 7 Review Answers
 - Highlighting and Note-Taking Modeling Chemistry Unit 7 Review Answers
 - Interactive Elements Modeling Chemistry Unit 7 Review Answers
8. Staying Engaged with Modeling Chemistry Unit 7 Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry Unit 7 Review Answers
9. Balancing eBooks and Physical Books Modeling Chemistry Unit 7 Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Chemistry Unit 7 Review Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Modeling Chemistry Unit 7 Review Answers
 - Setting Reading Goals Modeling Chemistry Unit 7 Review Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modeling Chemistry Unit 7 Review Answers
 - Fact-Checking eBook Content of Modeling Chemistry Unit 7 Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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