

Name _____

Chemistry Period _____

Worksheet Specific Heat

Specific Heat of water = $4.18 \text{ J/g}^\circ\text{C}$

1. A sample of mercury is heated from 25.5°C to 52.5°C . In the process, 3050J of heat are absorbed. What mass of mercury was in the sample? The specific heat of mercury is $0.140 \text{ J/g}^\circ\text{C}$. **ANS: 807 g**
2. A block of aluminum with a mass of 140g is cooled from 98.4°C to 62.2°C with a release of 1137J of heat. From these data, calculate the specific heat of aluminum. **ANS: 0.224**
3. A total of 226 J of heat are absorbed as 58.3 g of lead is heated from 12.0°C to 42.0°C . From this data, what is the specific heat of lead? **ANS: 0.129**
4. What is the change in heat energy when 132g of copper at 32.2°C is raised to 45.0°C ? The specific heat of copper is $0.385 \text{ J/g}^\circ\text{C}$. **ANS: 650 J**
5. What is the change in heat energy when 40.0g of water at 25.0°C is raised to 60.0°C ? **ANS: 5852 J**
6. What is the change in heat energy when 75.0 of water at 9.8°C is raised to 22.4°C ? **3950 J**
7. What is the change in heat energy when 125g of water at 10.0°C is raised to 90.0°C ? **ANS: 41.8 kJ**
8. What is the change in heat energy when 64.82 g of aluminum metal at 100.0°C is cooled to 82.0°C ? The specific heat of aluminum is $0.897 \text{ J/g}^\circ\text{C}$. **ANS: 1047**
9. What is the change in temperature if 24g of water at 50°C absorbs 985 J of heat energy? **ANS: 9.82**
10. What is the mass of a sample of iron if that sample lost 2300J of heat energy when it cooled from 80°C to 30°C ? The specific heat of iron is $0.449 \text{ J/g}^\circ\text{C}$. **ANS: 102 g**
11. What is the specific heat of metal if its mass is 26.86g and it requires 418.6 J of heat energy to raise its temperature from 27.4°C to 67.3°C ? **ANS: 0.391**
12. How much heat energy is needed to raise the temperature of 59.7g of cadmium from 25°C to 100°C ? The specific heat of cadmium is $0.231 \text{ J/g}^\circ\text{C}$. **ANS: 1034 J**

Math Skills Specific Heat Answers

Jianjun Gao



Math Skills Specific Heat Answers:

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Table of Contents Math Skills Specific Heat Answers

1. Understanding the eBook Math Skills Specific Heat Answers
 - The Rise of Digital Reading Math Skills Specific Heat Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Math Skills Specific Heat Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Math Skills Specific Heat Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Math Skills Specific Heat Answers
 - Personalized Recommendations
 - Math Skills Specific Heat Answers User Reviews and Ratings
 - Math Skills Specific Heat Answers and Bestseller Lists
5. Accessing Math Skills Specific Heat Answers Free and Paid eBooks

- Math Skills Specific Heat Answers Public Domain eBooks
 - Math Skills Specific Heat Answers eBook Subscription Services
 - Math Skills Specific Heat Answers Budget-Friendly Options
6. Navigating Math Skills Specific Heat Answers eBook Formats
- ePub, PDF, MOBI, and More
 - Math Skills Specific Heat Answers Compatibility with Devices
 - Math Skills Specific Heat Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Math Skills Specific Heat Answers
 - Highlighting and Note-Taking Math Skills Specific Heat Answers
 - Interactive Elements Math Skills Specific Heat Answers
8. Staying Engaged with Math Skills Specific Heat Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Math Skills Specific Heat Answers
9. Balancing eBooks and Physical Books Math Skills Specific Heat Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Math Skills Specific Heat Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Math Skills Specific Heat Answers
- Setting Reading Goals Math Skills Specific Heat Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Math Skills Specific Heat Answers
- Fact-Checking eBook Content of Math Skills Specific Heat Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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