

Assessment

Ions in Aqueous Solutions and Colligative Properties

Section Quiz: Colligative Properties of Solutions

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- B 1. Why is freezing-point depression a colligative property?
- It is inversely proportional to the number of particles in a solution.
 - It is directly proportional to the number of particles in a solution.
 - It depends on the properties of an electrolyte in a solvent.
 - None of the above.

- B 2. Compared with a 0.01 M $C_6H_{12}O_6$ solution, a 0.01 M KCl solution has
- the same freezing-point depression.
 - about twice the freezing-point depression.
 - the same freezing-point elevation.
 - about six times the freezing-point elevation.
- Handwritten notes: $C_6H_{12}O_6$ $i=1$, KCl $i=2$*

- D 3. Compared with a 1.00 M NaI solution, a 1.00 M Na_2SO_4 solution has
- the same boiling-point elevation.
 - about twice the boiling-point elevation.
 - a boiling-point elevation about two-thirds as high.
 - a boiling-point elevation about 1.5 times as high.
- Handwritten notes: NaI $i=2$, $Na_2(SO_4)$ $i=3$*

- B 4. When a nonvolatile solute dissolves in a solvent, the vapor pressure of the solvent
- increases.
 - decreases.
 - stays the same.
 - changes depending on the solvent used.
- Handwritten note: vapor pressure lowering (colligative property)*

- C 5. Which of the following compounds would be most effective in lowering the melting point of ice on roads?

a. $CaCl_2$

b. NaCl

c. K_3PO_4

d. K_2SO_4



$i=3$



$i=2$



$i=4$



$i=3$

Math Ch 13 Quiz 2 Answer Key

Raffaela Di Napoli



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